

CHAPTER 8

BERG/STRYER 8e STUDY GUIDE

1. Homework 1-4, 29, 35, 37, 38. Understand the general statements about enzymes on 215- 225: they are very **specific**; they can be **regulated** several ways (feedback/allosteric inhibition, regulatory proteins, covalent modification, and zymogen cleavage, as examples); they can **transform** energy; they **can't** move the **equilibrium** point; they stabilize the **transition state**; they have an **active site cleft**; and they often **change shape** in response to the "correct" substrate (**Koshland's Induced Fit**, Fig. 8.9). Note – 6 Enzyme classes "OTHLIL" Table 8.8 p. 245.

2. An important point in the chapter is the discussion of **Gibbs' Free Energy**, or ΔG . At equilibrium, the **enthalpy** change for a reaction will equal the **entropy** change times the absolute temperature, that is: $\Delta H = T\Delta S$. For reactions *not* at equilibrium, these two quantities differ, and the difference shows the available reaction energy, thus: $\Delta G = \Delta H - T\Delta S$. Defined this way – reactions *at* equilibrium have $\Delta G = 0$. Reactions which will move forward spontaneously have $\Delta G < 0$, i.e. negative. Reactions which will back up have $\Delta G > 0$. When standard conditions are applied, that is, all concentrations are set to [1 M], then the value of the free energy change will be ΔG° , the **Standard Free Energy Change**. Know the 5 statements listed on p. 218, and Equations (1) and (5) on 219. Be sure you understand Table 8.3. A difference in ΔG° of roughly 1.4 kcal/mol corresponds to a tenfold difference in the equilibrium constant. Thus, linking ATP hydrolysis to some reaction, which brings in -7.3 kcal/mol, ought to make something like a million-fold difference in the combined K_{eq} .

3. The single most important concept in this chapter is the **Michaelis Menten Equation**, p. 228 (eq. 26). You should memorize the equation in this form and be able to use it. You should also *understand* the derivation on 226-8. **Rate constants** (lowercase **k**) only "look" at a reaction in a single direction, whereas equilibrium constants (capital **K**) "look" at both directions. Rate constants fall out simply from the *Principle of Mass Action*. You should be able to figure out the dimensions (units) of the various rate constants in the M-M derivation. You should also know the underlying assumptions – mainly, "**Initial Rate**" (i.e. $[P] = 0$) and "**Steady State**" ($[ES]$ is constant, so $V_{fwd} = V_{back}$). You should also understand that the Michaelis Menten Equation assumes a single active site and a single substrate, so that many enzymes fit *other* kinetic patterns. The Michaelis-Menten equation graphs as a *rectangular hyperbola* (like Myoglobin's oxygen binding curve). **Allosteric** enzymes, usually with multiple interacting active sites, graph as *sigmoidal* plots (Fig 8.13 p. 233, see also Ch. 10 Fig 10.14). Understand the **double reciprocal plot** shown in Fig. 8.12 (p. 229) but don't use equation (eq. 27) please.

4. Understand the significance of **K_m** , **V_{max}** and turnover number **k_2** (p. 229 ff). Notice that the K_m varies in Table 8.4 from 0.4 to 8000 μM . A "typical" K_m might be about 1 mM or 1000 μM . Discussing a "typical" V_{max} is rather meaningless since $V_{max} = k_2 [E_t]$, and the enzyme concentration can be nearly anything. Instead, the "**turnover number**" or **k_2** is often discussed – in Table 8.5, it varies from 0.5 to 600,000 per second. Once again, a "typical" number might be about 300/sec. A high turnover number means a "fast" enzyme. A low K_m means an enzyme binds tightly and specifically to its substrate. The term " **k_{cat}** " (same as k_2 , turnover number) is defined on 230, and mainly is used to discuss the maximum rates of more complicated, non-M- M enzymes. Be able to distinguish between **competitive**, **uncompetitive**, and **non-competitive** inhibition (p. 234 ff).

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Table of Contents Berg Stryer 7e Study Guide Rutgers University

1. Understanding the eBook Berg Stryer 7e Study Guide Rutgers University
 - The Rise of Digital Reading Berg Stryer 7e Study Guide Rutgers University
 - Advantages of eBooks Over Traditional Books
2. Identifying Berg Stryer 7e Study Guide Rutgers University
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Berg Stryer 7e Study Guide Rutgers University
 - User-Friendly Interface
4. Exploring eBook Recommendations from Berg Stryer 7e Study Guide Rutgers University

- Personalized Recommendations
- Berg Stryer 7e Study Guide Rutgers University User Reviews and Ratings
- Berg Stryer 7e Study Guide Rutgers University and Bestseller Lists
- 5. Accessing Berg Stryer 7e Study Guide Rutgers University Free and Paid eBooks
 - Berg Stryer 7e Study Guide Rutgers University Public Domain eBooks
 - Berg Stryer 7e Study Guide Rutgers University eBook Subscription Services
 - Berg Stryer 7e Study Guide Rutgers University Budget-Friendly Options
- 6. Navigating Berg Stryer 7e Study Guide Rutgers University eBook Formats
 - ePub, PDF, MOBI, and More
 - Berg Stryer 7e Study Guide Rutgers University Compatibility with Devices
 - Berg Stryer 7e Study Guide Rutgers University Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Berg Stryer 7e Study Guide Rutgers University
 - Highlighting and Note-Taking Berg Stryer 7e Study Guide Rutgers University
 - Interactive Elements Berg Stryer 7e Study Guide Rutgers University
- 8. Staying Engaged with Berg Stryer 7e Study Guide Rutgers University
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Berg Stryer 7e Study Guide Rutgers University
- 9. Balancing eBooks and Physical Books Berg Stryer 7e Study Guide Rutgers University
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Berg Stryer 7e Study Guide Rutgers University
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Berg Stryer 7e Study Guide Rutgers University
 - Setting Reading Goals Berg Stryer 7e Study Guide Rutgers University
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Berg Stryer 7e Study Guide Rutgers University

- Fact-Checking eBook Content of Berg Stryer 7e Study Guide Rutgers University
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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