

Questions & Answers on Coordination Chemistry (D Ray)

1. What are the geometries of the following two complexes?

(i) $[\text{AlCl}_4]^-$

(ii) $[\text{Ag}(\text{NH}_3)_2]^+$

Ans. tetrahedral, linear

2. What is the respective central-metal oxidation state, coordination number, and the overall charge on the complex ion in $\text{NH}_4[\text{Cr}(\text{NH}_3)_4(\text{NCS})_4]$?

Ans. +3; 6; 1-

3. The formula for lithium iodotris(trifluorophosphine)nickelate(0) is: (Note: trifluorophosphine is a neutral ligand with the formula PF_3 .)

Ans. $\text{Li}[\text{Ni}(\text{PF}_3)_3\text{I}_3]$

4. Which is more likely to form a high-spin complex—*en*, F^- , or CN^- ?

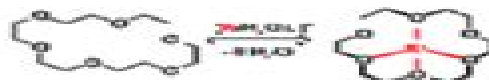
Ans. F^-

5. Name the following compound: $\text{K}_3[\text{CrCO}(\text{CN})_5]$.

Ans. potassium carbonylpentacyanochromium(III)

6. What is macrocyclic effect?

Ans. Macrocyclic ligands of appropriate size form more stable complexes than chelate ligands. A competition between a noncyclic chelating ligand and a macrocyclic (chelating) ligand having the same number and type of donor atoms will generally lead to complex formation predominantly by the macrocyclic ligand. This is known as the macrocyclic effect.



The entropy term drives this type of reaction to the right. The long noncyclic ligand is more flexible than the corresponding macrocyclic ligand and can adopt many more conformations than the macrocyclic ligand when it is not coordinated.

7. Predict the no. of unpaired electrons, the spin-only magnetic moments at 25 °C for

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Richard L. Middaugh



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