

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Covalent Bonding Chapter 8 Worksheet Answers

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In a world inundated with displays and the cacophony of instantaneous connection, the profound energy and emotional resonance of verbal beauty frequently fade in to obscurity, eclipsed by the continuous barrage of noise and distractions. Yet, located within the lyrical pages of **Covalent Bonding Chapter 8 Worksheet Answers Alitaoore**, a charming work of fictional beauty that pulses with fresh feelings, lies an remarkable trip waiting to be embarked upon. Published by way of a virtuoso wordsmith, this mesmerizing opus books readers on a psychological odyssey, delicately revealing the latent potential and profound influence embedded within the delicate web of language. Within the heart-wrenching expanse of this evocative analysis, we shall embark upon an introspective exploration of the book is central themes, dissect its captivating writing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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