

\*List given in descending order of diameter.

As charges go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing on the top. It's most organized alphabetically as elements can be found easily. But as a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. Those properties relate to the electronic structure of the atoms of each element. In today's activities you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

100

1				$\frac{1}{2}$	$\frac{1}{2}$	$\frac{1}{2}$					$\frac{1}{2}$	$\frac{1}{2}$	$\frac{1}{2}$	
2				$\frac{1}{2}$	$\frac{1}{2}$	$\frac{1}{2}$					$\frac{1}{2}$	$\frac{1}{2}$	$\frac{1}{2}$	

- | $XY$             | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ |
|------------------|------------------|------------------|------------------|------------------|
| $Y_{\text{obs}}$ | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ | $Y_{\text{obs}}$ |

- Answers will vary for each group. For Team 1, the configuration should be as follows:

# Cracking The Periodic Table Code Answers

**Thomas Griffiths**



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