

CHEMCOMM
Lab, Molar Heat of Combustion of a Candle

NAME: _____
PERIOD: _____

Purpose: To experimentally determine the heat of combustion of a candle

Introduction: During chemical reactions, both matter and energy are conserved. In other words, even though they are rearranged during the course of the reaction, both are constant in the universe. The combustion of wax is the reaction of the *hydrocarbons* ($\sim C_{25}H_{52}$) *with oxygen in the atmosphere to produce carbon dioxide and water vapor*. This reaction is exothermic, i.e. releases heat to the surroundings. The amount of energy released can be measured through by using the calorimetry procedure.

Materials: Candle in holder
Ring stand with ring
Stirring rod
50-ml graduated cylinder
Thermometer
Empty soda can with tab top

Safety:

- Tie back hair, loose clothing and jewelry
- Handle lighted candle with care; wax will be hot after burning
- Dispose of matches in according to teacher instructions

Procedure:

1. Use the following table to collect the data taken throughout the procedure:

Initial mass of candle	
Final mass of candle	
Volume of water in can	
Room temperature	
Initial temperature of water in can	
Final temperature of water in can	

2. Take the mass of the candle in its holder and record in the data table.
3. Exactly measure 100ml of chilled water supplied by your teacher. The volume doesn't really matter, but knowing the exact amount you have does. Pour the water into your empty soda can.
4. Set-up calorimetry apparatus as shown by your teacher.
5. Measure and record the room temperature (i.e. just read the temperature on your thermometer)
6. Measure and record the temperature of the water in the soda can.
7. Light the candle and place it under your soda can so that the flame is just below the bottom of the can. Gently stir the water with your other stirring rod as it heats. Take care not to break the stirring rod.
8. Continue heating the water until the temperature rises as far above room temperature as it was below room temperature to start.

Heat Combustion Candle Lab Answers

M Lipman



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