

15% solution by mass means

Mass of solution = 100g

Mass of solute = 15g

Number of moles of  $\text{H}_2\text{SO}_4 = 15/98$

$\text{H}_2\text{SO}_4$

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}}$$

$$1 \times 2 + 32 + 16 \times 4 = 98$$

$$1.2 = \frac{100}{\text{Volume}}$$

$$\text{Volume} = \frac{100}{2} \times 10 = \frac{1000}{12} = 0.012 \text{ ml}$$

$$\frac{100}{1.2} \text{ solution contain} = \frac{15}{98}$$

$$1 \text{ solution contain} = \frac{15}{98} \times \frac{10}{100}$$

$$1000 \text{ solution contain} = \frac{15}{98} \times \frac{1.2}{100} \times 1000$$

$$= 1.8 \text{ M.}$$

#### 2.4|MOLALITY

Number of moles of the solute present in 1000 g solvent. It is represented by 'M'.

#### 2.5|NORMALITY

Number of gram equivalents of a solute in 1000 ml solution. It is represented by 'N'.

$$\text{Number of gram equivalents} = \frac{\text{Mass in gram}}{\text{equivalent Mass}}$$

#### FORMULAE FOR EQUIVALENT MASS

$$\text{Equivalent mass of acid} = \frac{\text{Molecular Mass of Acid}}{\text{Basicity}}$$

$$\text{Equivalent mass of base} = \frac{\text{Molecular Mass of Base}}{\text{Acidity}}$$

$$\text{Equivalent mass of salt} = \frac{\text{Formula mass of salt}}{\text{tvevalency of cation}}$$

#### 2.6|MOLE FRACTION

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