

LAB 06 – DOUBLE DISPLACEMENT REACTIONS

Discussion - Double Replacement Reactions (Refer to class notes from Chapter 7)

All reactions in this part of the experiment are double replacement reactions.

Double replacement reactions, sometimes called "double displacement", "exchange reactions" or "trading partners" are very common. Generally, they occur when ionic compounds are dissolved in aqueous (water) solution. They have the general form:



where A and B are positive ions and X and Y are negative ions.

A reaction will take place if one of the following is formed in the reaction:

- (a) a **precipitate** (what you see is a solid material that clouds the solution).
- (b) a **gas** (you see lots of bubbles coming out of solution).
- (c) **water** (You will feel the test tube get warm as heat is given off by the reaction).

These are known as the **Driving Force** of the reaction.

If the reaction produces neither a precipitate, nor a gas, nor a molecular substance, then **no reaction** has taken place. If a reaction does take place, then moving from a balanced molecular equation to a net ionic equation allows us to determine exactly what the chemistry was in the reaction, eliminating the non-reacting (spectator) ions and/or molecules.

The following double displacement reactions were carried out in lab and the listed observations were made.

Reagents	Observations
1. $\text{NaCl(aq)} + \text{AgNO}_3\text{(aq)}$	<i>White precipitate formed</i>
2. $\text{Na}_2\text{CO}_3\text{(aq)} + \text{HCl(aq)}$	<i>Bubbles form, solution got warm</i>
3. $\text{NaOH(aq)} + \text{HCl(aq)}$	<i>Solution got warm</i>
4. $\text{Na}_2\text{CO}_3\text{(aq)} + \text{CaCl}_2\text{(aq)}$	<i>White precipitate formed</i>
5. $\text{FeCl}_3\text{(aq)} + \text{NH}_4\text{OH(aq)}$	<i>Reddish brown precipitate formed</i>
6. $\text{CuSO}_4\text{(aq)} + \text{Zn(NO}_3)_2\text{(aq)}$	<i>No Reaction</i>

Double Displacement Reaction Lab Answers

RS Peters



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Lab Experiments in Introductory Chemistry Phil Reedy, Donald J. Wink, Sharon Fetzer-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at <http://custompub.whfreeman.com>

Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001

Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

Foundations of College Chemistry Morris Hein, Susan Arena, Cary Willard, 2016-08-02 This text is an unbound three hole punched version Used by over 750 000 students Foundations of College Chemistry Binder Ready Version 15th Edition is praised for its accuracy clear no nonsense approach and direct writing style Foundations direct and straightforward explanations focus on problem solving making it the most dependable text on the market Its comprehensive scope proven track record outstanding in text examples and problem sets were all designed to provide instructors with a solid text while not overwhelming students in a difficult course Foundations fits into the prep intro chemistry courses which often include a wide mix of students from science majors not yet ready for general chemistry allied health students in their 1st semester of a GOB sequence science education students for elementary school teachers to the occasional liberal arts student fulfilling a science requirement Foundations was specifically designed to meet this wide array of needs

Inquiry-based Experiments in Chemistry Valerie Ludwig Lechtanski, 2000 Inquiry Based Experiments in Chemistry is an alternative to those cookbook style lab manuals providing a more accurate and realistic experience of scientific investigation and thought for the high school chemistry or physical science student

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

Carolina Science and Math Carolina Biological Supply Company, 2003

Steps to Doing Science Kristin Tuttle Bump, 2008

Prentice Hall Chemistry, 2000

Authentic Practice Chemistry David G. White, 2000

Chemistry McGraw-Hill Staff, 2001-07

School Library Journal, 1986

Prentice Hall Science Explorer Michael J. Padilla, 2002

Student Study Guide to accompany Chemistry Martin Silberberg, Libby Weberg, 2005-01-06

Fundamentals of Chemistry in the Laboratory Ralph A. Burns, 1999

Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence

with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students **General, Organic, and Biological Chemistry Study Guide and Selected Solutions** Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems Laboratory Exercises for Preparatory Chemistry Kathy Dodds Tyner, 1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course Tyner s manual emphasizes the application of chemistry and the principles of science to everyday life The labs are directly applicable to the real world and often contain supplemental assignments that illustrate an application *Essentials of Introductory Chemistry* Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry *Energy Research Abstracts* , 1987

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