- According to the Bronsted—Lowry definition, which chemical species can function both as an acid and as a base?
  - a. CI = 6. SO<sub>2</sub> =
  - SO/\*\*
     NH/\*
  - d. HCO<sub>2</sub>
  - e. H<sub>2</sub>O\*
- 2. In the reaction CN<sup>-</sup> + H<sub>2</sub>O ← → HCN + OH<sup>-</sup> which is an acid—base conjugate pair?
  - a. H<sub>2</sub>O and HCN
  - b. H<sub>2</sub>O and OH<sup>-</sup>
  - c. ON and H<sub>2</sub>O
  - d. HCN and OH"
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX
  - a. It is less basic than a I M solution of NaY.
  - b. It is more basic than a I M solution of NaY.
  - It yields a neutral solution.
  - It is more concentrated than a 1 M solution of NaY.
- 4. The pH of a 0.03 M HCI solution is
  - a. 1.5
  - 6. 25
  - c. 3.5
  - d. 12.5
- The pH of a 1.0 x 10<sup>-3</sup> M Ba(OH)<sub>2</sub> solution at 25°C is
  - a. 2.7
  - b. 3.0
  - c. 11.0
  - d. 11.3
- 6. Which is the strongest acid?
  - a. HCIO
  - A. HCIO<sub>2</sub>
  - c. HOIO<sub>3</sub>
  - d. HClO<sub>4</sub>
- 7. Which salt reacts with water (hydrolyzes) to produce a basic solution?
  - a. NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
  - NaNO<sub>3</sub>
  - c. NH<sub>f</sub>CI
  - d. BaSO<sub>4</sub>

8. The weakest of the bases listed is

Acid	Conjugate Base	Ka (Ionization Constant of Acid)				
HCL	CIT	100% ionized				
HSO <sub>4</sub>	SO <sub>4</sub> 2-	1.2 x 10°				
H <sub>2</sub> S HS <sup>-</sup> HS <sup>-</sup> S <sup>r-</sup>		5.7x10 *				
		1.2x10 <sup>-0</sup>				

- a. CIT
- # 1.00 T
- e. HS<sup>-</sup> d. S<sup>3-</sup>
- e. 50e<sup>2-</sup>
- 9. The oxide of which element will react with water to form the strongest acid?

Period	Main Groups							
	1	П	П	IV	V	VI	VII	0
First								
Second	X	Y	Z	P		Q	S	U
Third	WAY					R	T	0.40

- e. W
- 6. M
- c. P
- d. R.
- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate KC<sub>2</sub>H<sub>2</sub>O<sub>2</sub>, is less alkaline than a 0.10 M solution of potassium evanide. KCN?
  - a. Hydrocyanic acid is a weaker acid than acetic acid.
  - Hydrocyanic acid is less soluble in water than acetic acid.
  - c. Cyanides are less soluble than acetaces
  - d. Acetic acid is a weaker acid than hydrocyanic acid
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of
  - a. a 0.100 M sodium solution.
  - a 0.0500 M sodium hydroxide solution.
  - c. a 0.0500 M benzoic acid solution.
  - a 0.0500 M sodium benzoate solution.

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