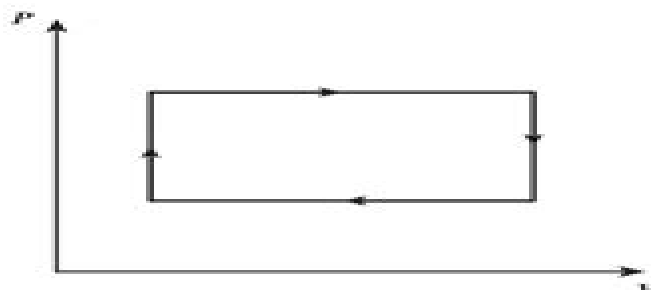


CHAPTER 9 PRACTICE QUESTIONS

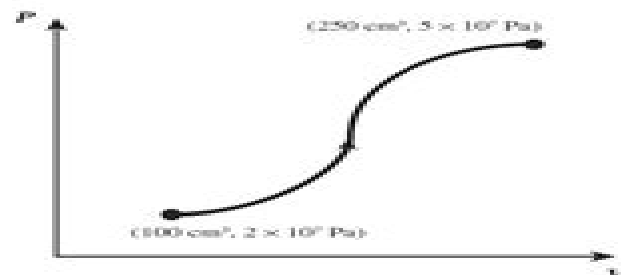
Directions: Complete the following problems as specified by each question, and then check your work using the solutions that follow. For extended, step-by-step solutions, access your Student Tools online.

1. Water has a specific heat of $4186 \text{ J/kg}\cdot\text{K}$. How much heat would it take to raise the temperature of 100 g of water from 25°C to 75°C ?
2. To change phase, a sample of a substance requires an amount of heat per unit mass known as the **latent heat** for that transformation. For evaporation, water has a latent heat of 2257 kJ/kg . How much heat is required to evaporate 30 g of water?
3. How much heat is required to completely evaporate 50 g of water at an initial temperature of 50°C ? Refer to the previous two problems for relevant constants.
4. If 3 mol of an atomic gas changes from a state $(0.1 \text{ m}^3, 1.0 \times 10^5 \text{ Pa})$ to $(0.2 \text{ m}^3, 3.0 \times 10^5 \text{ Pa})$ due to some arbitrary process, and absorbs 45 kJ of heat in the process, how much work is done? Is this an example of an engine?

5. Does the gas in the following cycle lose or gain heat?



6. What is the change in internal energy for a sample of molecular hydrogen undergoing the following process? Assume that the hydrogen atoms are connected rigidly.



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Table of Contents High School Physics Questions And Answers

1. Understanding the eBook High School Physics Questions And Answers
 - The Rise of Digital Reading High School Physics Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying High School Physics Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High School Physics Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from High School Physics Questions And Answers
 - Personalized Recommendations
 - High School Physics Questions And Answers User Reviews and Ratings

- High School Physics Questions And Answers and Bestseller Lists
- 5. Accessing High School Physics Questions And Answers Free and Paid eBooks
 - High School Physics Questions And Answers Public Domain eBooks
 - High School Physics Questions And Answers eBook Subscription Services
 - High School Physics Questions And Answers Budget-Friendly Options
- 6. Navigating High School Physics Questions And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - High School Physics Questions And Answers Compatibility with Devices
 - High School Physics Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of High School Physics Questions And Answers
 - Highlighting and Note-Taking High School Physics Questions And Answers
 - Interactive Elements High School Physics Questions And Answers
- 8. Staying Engaged with High School Physics Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High School Physics Questions And Answers
- 9. Balancing eBooks and Physical Books High School Physics Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection High School Physics Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine High School Physics Questions And Answers
 - Setting Reading Goals High School Physics Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of High School Physics Questions And Answers
 - Fact-Checking eBook Content of High School Physics Questions And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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