

## Alka Seltzer Strength

### Lab Report

This experiment explores the chemistry behind common antacids and treatment for heartburn. This lab continues to build on the concept of stoichiometry and limiting reactant explored in previous labs. It also uses the Law of Conservation of Mass to quantitatively (with numbers and measurements) study the amount of substance used and produced in a chemical reaction.

After performing the lab, finding the limiting reactant, and determining the maximum amount of carbon dioxide produced, you will be able to calculate the mass percent of the active ingredient in a tablet of Alka Seltzer.

### Data and Calculations

Data 6 points

Beaker	Mass Empty Beaker (g)	Mass of Beaker w/water and vinegar (g)	Mass before reaction; Beaker, water, vinegar and 3.000 g Alka Seltzer (g)	Mass after reaction; beaker and contents after reaction is complete(g)
1	32.000 g	62.000 g	65.000 g	64.239 g
2	32.000 g	62.005 g	65.005 g	64.171 g
3	32.000 g	62.010 g	65.010 g	64.127 g
4	32.000 g	62.015 g	65.015 g	64.132 g
5	32.000 g	62.020 g	65.020 g	64.137 g
6	32.000 g	62.026 g	65.026 g	64.142 g

### Calculations

12 points

Beaker	Change in Mass (grams of CO <sub>2</sub> made)	Moles of CO <sub>2</sub> made (Molar mass of CO <sub>2</sub> is 44.01 g/mol)	Moles of NaHCO <sub>3</sub> reacted	Grams of NaHCO <sub>3</sub> reacted (molar mass of NaHCO <sub>3</sub> is 84.007g/mol)
1				
2				
3				
4				
5				
6				

# Alka Seltzer Lab Answers

**Keith Johnson**



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