

Analysis of a Mixture of Carbonate and Bicarbonate

The total alkalinity (defined as total base concentration in this case $[HCO_3^-] + 2[CO_3^{2-}]$) is measured by titrating the mixture with standard HCl to a bromocresol green end point:



Assuming all the carbonate comes from sodium carbonate and all the bicarbonate comes from sodium bicarbonate, calculate the original percent composition of these two substances in your unknown mixture.

REAGENTS

Standard 0.1 M HCl (note you must standardize it yourself):

CO₂-free water: Boil ~600 mL of distilled water to expel CO₂ and pour the water into a 1000-mL plastic bottle. Screw the cap on tightly and allow the water to cool to room temperature. Keep tightly capped when not in use.

Bromocresol green. (See Table 11-4 in the textbook.)

Unknowns: Solid unknowns (2.5 g/student) prepared from reagent-grade sodium carbonate and sodium bicarbonate. Unknowns should be stored in a desiccator and should not be heated. Heating at 50°C–100°C converts NaHCO₃ to Na₂CO₃.

Sodium carbonate (ACS grade) ~2 g/student for standardization of HCl

PROCEDURE

1. Accurately weigh 2.0–2.5 g of Sodium carbonate standard into a 250-mL volumetric flask by weighing the sample in a capped weighing bottle, delivering some to a funnel in the volumetric flask, and reweighing the bottle. Continue this process until the desired mass of reagent has been transferred to the funnel. Rinse the funnel repeatedly with small portions of CO₂-free water to dissolve the sample. Remove the funnel, dilute to the mark, and mix well (with CO₂ free water). Pipet a 25.00-mL aliquot of standard solution into a 250-mL flask and titrate with ~0.1 M HCl, using bromocresol green indicator. Repeat this procedure with three more 25.00-mL aliquots. Express your Molarity of your standardized HCl as $\bar{x} \pm s$.

2. Accurately weigh 2.0–2.5 g of unknown into a 250-mL volumetric flask by weighing the sample in a capped weighing bottle, delivering some to a funnel in the volumetric flask, and reweighing the bottle. Continue this process until the desired mass of reagent has been transferred to the funnel. Rinse the funnel repeatedly with small portions of CO₂-free water to dissolve the sample. Remove the funnel, dilute to the mark, and mix well (with CO₂ free water).

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American Gas Institute



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Calculations of Quantitative Chemical Analysis Leicester Forsyth Hamilton, Stephen Gershom Simpson, 1922

Technical Paper - Bureau of Mines United States. Bureau of Mines, 1927 *Manual of American Steel & Wire Company's Process of Water Purification with Sulphate of Iron ...* American Steel & Wire Co, 1916 The Journal of Industrial and Engineering Chemistry, 1923 **Industrial & Engineering Chemistry**, 1923 **Evaluation of Streamflow, Water Quality, and Permitted and Nonpermitted Loads and Yields in the Raritan River Basin, New Jersey, Water Years 1991-98** Robert G. Reiser, 2004 *Water-resources Investigations Report* Robert G. Reiser, 2004 Water-resources Investigations Report, 2003 **Quantitative Analysis** Stephen Popoff, 1927 Environmental Sampling and Analysis Maria Csuros, 2018-05-11 This manual covers the latest laboratory techniques state of the art instrumentation laboratory safety and quality assurance and quality control requirements In addition to complete coverage of laboratory techniques it also provides an introduction to the inorganic nonmetallic constituents in environmental samples their chemistry and their control by regulations and standards Environmental Sampling and Analysis Laboratory Manual is perfect for college and graduate students learning laboratory practices as well as consultants and regulators who make evaluations and quality control decisions Anyone performing laboratory procedures in an environmental lab will appreciate this unique and valuable text *Chemical Abstracts*, 1918 **Proceedings of National Electric Light Association** National Electric Light Association. Convention, 1929 **Report of Prime Movers Committee, Technical National Section ... Presented at the ... Convention of the National Electric Light Association** National Electric Light Association. Prime Movers Committee, 1929 Circular Pennsylvania State University. Earth and Mineral Sciences Experiment Station, 1929 Cheap Steam, 1925 Boiler Feed Water Purification Sheppard T. Powell, 1927 **Proceedings** Brian Hitchon, Stefan Bachu, Claude M. Sauveplane, 1986 **Official Proceedings**, 1924 **Proceedings of the American Gas Institute** American Gas Institute, 1917 List of members in v 2 4 11 *Journal of Gas Lighting and Water Supply*, 1926

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