

## CHAPTER 8

### BERG/STRYER 8e STUDY GUIDE

1. Homework 1-4, 29, 35, 37, 38. Understand the general statements about enzymes on 215- 225: they are very **specific**; they can be **regulated** several ways (feedback/allosteric inhibition, regulatory proteins, covalent modification, and zymogen cleavage, as examples); they can **transform** energy; they **can't** move the **equilibrium** point; they stabilize the **transition state**; they have an **active site cleft**; and they often **change shape** in response to the "correct" substrate (**Koshland's Induced Fit**, Fig. 8.9). Note – 6 Enzyme classes "OTHLIL" Table 8.8 p. 245.

2. An important point in the chapter is the discussion of **Gibbs' Free Energy**, or  $\Delta G$ . At equilibrium, the **enthalpy** change for a reaction will equal the **entropy** change times the absolute temperature, that is:  $\Delta H = T\Delta S$ . For reactions *not* at equilibrium, these two quantities differ, and the difference shows the available reaction energy, thus:  $\Delta G = \Delta H - T\Delta S$ . Defined this way – reactions *at* equilibrium have  $\Delta G = 0$ . Reactions which will move forward spontaneously have  $\Delta G < 0$ , i.e. negative. Reactions which will back up have  $\Delta G > 0$ . When standard conditions are applied, that is, all concentrations are set to [1 M], then the value of the free energy change will be  $\Delta G^\circ$ , the **Standard Free Energy Change**. Know the 5 statements listed on p. 218, and Equations (1) and (5) on 219. Be sure you understand Table 8.3. A difference in  $\Delta G^\circ$  of roughly 1.4 kcal/mol corresponds to a tenfold difference in the equilibrium constant. Thus, linking ATP hydrolysis to some reaction, which brings in -7.3 kcal/mol, ought to make something like a million-fold difference in the combined  $K_{eq}$ .

3. The single most important concept in this chapter is the **Michaelis Menten Equation**, p. 228 (eq. 26). You should memorize the equation in this form and be able to use it. You should also *understand* the derivation on 226-8. **Rate constants** (lowercase **k**) only "look" at a reaction in a single direction, whereas equilibrium constants (capital **K**) "look" at both directions. Rate constants fall out simply from the *Principle of Mass Action*. You should be able to figure out the dimensions (units) of the various rate constants in the M-M derivation. You should also know the underlying assumptions – mainly, "**Initial Rate**" (i.e.  $[P] = 0$ ) and "**Steady State**" ( $[ES]$  is constant, so  $V_{fwd} = V_{back}$ ). You should also understand that the Michaelis Menten Equation assumes a single active site and a single substrate, so that many enzymes fit *other* kinetic patterns. The Michaelis-Menten equation graphs as a *rectangular hyperbola* (like Myoglobin's oxygen binding curve). **Allosteric** enzymes, usually with multiple interacting active sites, graph as *sigmoidal* plots (Fig 8.13 p. 233, see also Ch. 10 Fig 10.14). Understand the **double reciprocal plot** shown in Fig. 8.12 (p. 229) but don't use equation (eq. 27) please.

4. Understand the significance of  **$K_m$** ,  **$V_{max}$**  and turnover number  **$k_2$**  (p. 229 ff). Notice that the  $K_m$  varies in Table 8.4 from 0.4 to 8000  $\mu M$ . A "typical"  $K_m$  might be about 1 mM or 1000  $\mu M$ . Discussing a "typical"  $V_{max}$  is rather meaningless since  $V_{max} = k_2 [E_t]$ , and the enzyme concentration can be nearly anything. Instead, the "**turnover number**" or  **$k_2$**  is often discussed – in Table 8.5, it varies from 0.5 to 600,000 per second. Once again, a "typical" number might be about 300/sec. A high turnover number means a "fast" enzyme. A low  $K_m$  means an enzyme binds tightly and specifically to its substrate. The term " **$k_{cat}$** " (same as  $k_2$ , turnover number) is defined on 230, and mainly is used to discuss the maximum rates of more complicated, non-M- M enzymes. Be able to distinguish between **competitive**, **uncompetitive**, and **non-competitive** inhibition (p. 234 ff).

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**D Keegan**



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web nov 11 2023 there are two ways in which the comparative and superlative forms of adjectives are formed you add er comparative or est superlative to the adjective adjectives with one syllable usually take these endings if the word already ends in e the e must be left off if a word ends in y it usually takes er or est and the y changes to i

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