

CHAPTER 9 REVIEW***Stoichiometry*****MIXED REVIEW****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Given the following equation:
- $\text{C}_3\text{H}_8(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

- _____ a. What is the value of the coefficient x in this equation?
- _____ b. What is the molar mass of C_3H_8 ?
- _____ c. What is the mole ratio of O_2 to H_2O in the above equation?
- _____ d. How many moles are in an 8.0 g sample of C_3H_8 ?
- _____ e. If z mol of C_3H_8 react, how many moles of CO_2 are produced, in terms of z ?

- 2.**
- a. What is meant by
- ideal conditions*
- relative to stoichiometric calculations?

- b. What function do ideal stoichiometric calculations serve?

- c. Are actual yields typically larger or smaller than theoretical yields?

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

- 3.**
- Assume the reaction represented by the following equation goes all the way to completion:



- _____ a. If 6 mol of H_2 are consumed, how many moles of NH_3 are produced?
- _____ b. How many grams are in a sample of NH_3 that contains 3.0×10^{23} molecules?

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