

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Student Exploration: Boyle's Law and Charles' Law

**Vocabulary:** absolute zero, Boyle's law, Charles' law, Gay-Lussac's law, Kelvin scale, pressure

**Prior Knowledge Question** (Do this BEFORE using the Gizmo.)

A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How can such a small tank contain enough helium to fill so many balloons?

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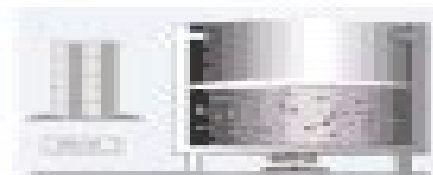
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### Gizmo Warm-up

The Boyle's Law and Charles' Law Gizmo™ shows a container of gas. In the container, the small purple spheres represent molecules.



1. Observe the particles. Are they all moving at the same speed? \_\_\_\_\_
2. How do the particles interact with the walls and lid of the container? \_\_\_\_\_

These interactions contribute to the ~~pressure~~ on the walls of the container. Pressure is defined as force per unit area. The SI units of pressure are newtons per square meter ( $\text{N/m}^2$ ), or pascals (Pa).

3. Slowly drag the temperature (T) slider back and forth. (Note: In this Gizmo, the **Kelvin scale** is used to measure temperature. On the Kelvin scale, 0 degrees is **absolute zero**, the coldest possible temperature. Absolute zero is equal to  $-273.15^\circ\text{C}$  or  $-459.67^\circ\text{F}$ .)
  - A. How does the change in temperature affect the speed of the molecules? \_\_\_\_\_
  - B. How does the change in temperature affect the volume of the container? \_\_\_\_\_

# Charles And Boyles Law Gizmo Answer Key Pdf

**D Siedentop**



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