# Chapter 10 mole concept advance additional notes

What is mole?



This is an atom of carbon-12. It is chosen as the standard atom. Its A, is taken as 12. Then other atoms are compared with it.



This is a magnesium atom. It is twice as heavy as a carbon-12 atom, so its A, is 24. So it follows that ...

A mole of a substance is defined as the number of atoms in an element More examples of moles

Sodium is made of single sodium atoms. Its symbol is Na. Its A, is 2.3.



todime is made of todime molecules. Its formula is I<sub>p</sub>. Its M, is 254.



Water is made of water molecules. Its formula is H<sub>i</sub>O. Its M<sub>i</sub> is 18.





This is 23 grams of sedium. It contains 6.02 × 10<sup>15</sup> sedium atoms, or 1 mole of sedium atoms.



This is 254 grams of todine. It contains  $6.02 \times 10^{2}$  indine molecules, or 1 mole of todine molecules.



The beaker contains 18 grams of water, or 8.02 × 10° water molecules, or 1 mole of water molecules.

#### Finding the mass of a mole

in order to find the mass of any substance, you must follow these steps;

- Write down the symbol or formula of the substance
- Find its M. or A.
- Express that mass in grams(g)

#### Examples:

Substance	Symbol or formula		AM,	Mass of 1 male	
hidaum	He	the - a	exists as single access.	4 grams	
cxygen	0,	0 - 16	2 × 10 × 32	32 grams	
ethanol	C,FC0H	C - 12 H - 1 O - 16	2 × 12 - 24 6 × 1 - 6 1 × 16 - 16 46	46 grams	

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