

## SOLVED QUESTION BANK

## Objective / Very Short Answer Type Questions [1 Mark]

1. The half life period of first order reaction is 1386 seconds. The specific rate constant of the reaction is

(a)  $0.5 \times 10^{-3} \text{ s}^{-1}$  (b)  $0.5 \times 10^{-3} \text{ s}^{-1}$   
(c)  $5.0 \times 10^{-2} \text{ s}^{-1}$  (d)  $5.0 \times 10^{-3} \text{ s}^{-1}$

Ans. (b)  $t_{1/2} = \frac{0.693}{k}$   
 $\Rightarrow k = \frac{0.693}{1386} = 0.5 \times 10^{-3} \text{ s}^{-1}$

2. The rate constant of a reaction  $A \longrightarrow B$  is  $0.6 \times 10^{-3}$  mole per second. If the concentration of [A] is 5 M, then what will be concentration of [B] after 20 minutes?

(a) 0.36 M (b) 0.72 M  
(c) 1.08 M (d) 3.60 M

Ans. (b)  $k = \frac{[R]_0 - [R]}{t}$  for zero order.  
 $\Rightarrow 0.6 \times 10^{-3} \times 20 \times 60 = [R]_0 - [R]$   
 $\Rightarrow 0.72 \text{ M}$   
 Amount of [B] produced =  $[R]_0 - [R] = 0.72 \text{ M}$

3. Which of the following is affected by catalyst?

(a)  $\Delta H$  (b)  $\Delta S$   
(c)  $\Delta G$  (d)  $E_a$

Ans. (d) Catalyst affects activation energy. It does not affect,  $\Delta H$ ,  $\Delta S$ ,  $\Delta G$ .

4. A first order reaction has specific reaction rate  $10^{-2} \text{ s}^{-1}$ . How much time it will take for 20g of reactant to reduce to 5g?

(a) 138.6 s (b) 346.5 s  
(c) 693.0 s (d) 238.6 s

Ans. (a)  $t_{3/4} = \frac{1.386}{k} = \frac{1.386}{10^{-2}} = 138.6 \text{ s}$

5. The activation energy of a reaction can be determined from the slope of which of the following graph:

(a)  $\ln k$  Vs  $\frac{1}{T}$  (b)  $\frac{T}{\ln k}$  Vs  $\frac{1}{T}$   
(c)  $\ln k$  Vs  $T$  (d)  $\frac{\ln k}{T}$  Vs  $T$

Ans. (a) When we plot  $\ln k$  Vs  $\frac{1}{T}$ , slope of line =  $-\frac{E_a}{R}$ .

6. Mechanism of a hypothetical reaction  $X_2 + Y_2 \longrightarrow 2XY$  is given below:

(a)  $X_2 \rightleftharpoons X + X$  (fast)  
(b)  $X + Y_2 \longrightarrow XY + Y$  (slow)  
(c)  $X + Y \longrightarrow XY$  (fast)

The overall order of reaction is

(a) 2 (b) 0 (c) 1.5 (d) 1

Ans. (c) rate =  $k[X][Y_2]$ ,  $K = \frac{[X]^2}{[X_2]} \Rightarrow [X] = \sqrt{K[X_2]}$   
 rate =  $k \times K^{1/2} [X_2]^{1/2} [Y_2] = k^1 [X_2]^{1/2} [Y_2]^1$   
 Order =  $\frac{1}{2} + 1 = \frac{3}{2} = 1.5$

7. The rate of first order reaction is  $0.04 \text{ mol L}^{-1} \text{ s}^{-1}$  at 10 sec, and  $0.03 \text{ mol L}^{-1}$  at 20 seconds after initiation of the reaction.  $t_{1/2}$  of reaction is

(a) 44.1 s (b) 54.1 s (c) 24.1 s (d) 34.1 s

Ans. (c)  $k = \frac{2.303}{20 - 10} \log \frac{0.04}{0.03} = \frac{2.303}{10} \log \frac{4}{3}$   
 $= \frac{2.303}{10} (\log 4 - \log 3)$   
 $= \frac{2.303}{10} (0.6021 - 0.4771) = \frac{2.303}{10} \times 0.125$   
 $t_{1/2} = \frac{0.693}{k} = \frac{0.693 \times 10}{2.303 \times 0.1249} = \frac{3.010}{0.1249} = 24.1 \text{ s}$

8. If the initial concentration of reactant is doubled  $t_{1/2}$  is also doubled, the order of reaction is

(a) zero (b) 1 (c) 2 (d) 3

Ans. (a) In zero order of reaction  $t_{1/2} \propto [R]_0$

9. For a reaction taking place in three steps, the overall rate constant  $k = \frac{k_1 k_2}{k_3}$ . If  $E_{a1}$ ,  $E_{a2}$  and  $E_{a3}$  are the

40, 50 and 60  $\text{kJ mol}^{-1}$ , the overall  $E_a$  is

(a) 30 (b) 40 (c) 60 (d) 50

Ans. (a)  $k = A e^{-E_a/RT}$   
 $k = \frac{k_1 k_2}{k_3}$   
 $\Rightarrow A e^{-E_a/RT} = \frac{A e^{-E_{a1}/RT} \times A e^{-E_{a2}/RT}}{A e^{-E_{a3}/RT}}$   
 $e^{-E_a/RT} = \frac{e^{-E_{a1}/RT} \times e^{-E_{a2}/RT}}{e^{-E_{a3}/RT}}$

Taking  $\ln$  on both side we get

$$\frac{E_a}{RT} = \frac{-E_{a1} - E_{a2} + E_{a3}}{RT}$$

$$E_a = E_{a1} + E_{a2} - E_{a3}$$

$$E_a = 40 + 50 - 60 = 30 \text{ kJ mol}^{-1}$$

10. If conc. of reactant 'A' is increased 10 times and rate of reaction becomes 100 times. What is order with respect to 'A'?

(a) 1 (b) 2 (c) 3 (d) 4

# Chemical Kinetics Questions And Answers

**Peter R. Bergethon, Kevin Hallock**



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*Chemical Kinetics and Mechanism* M Mortimer, P G Taylor, 2007-10-31 Chemical Kinetics and Mechanism considers the role of rate of reaction It begins by introducing chemical kinetics and the analysis of reaction mechanism from basic well established concepts to leading edge research Organic reaction mechanisms are then discussed encompassing curly arrows nucleophilic substitution and E1 and E2 elimination reactions The book concludes with a Case Study on Zeolites which examines their structure and internal dimensions in relation to their behaviour as molecular sieves and catalysts The accompanying CD ROM contains the Kinetics Toolkit a graph plotting application designed for manipulation and analysis of kinetic data which is built into many of the examples questions and exercises in the text There are also interactive activities illustrating reaction mechanisms The Molecular World series provides an integrated introduction to all branches of chemistry for both students wishing to specialise and those wishing to gain a broad understanding of chemistry and its relevance to the everyday world and to other areas of science The books with their Case Studies and accompanying multi media interactive CD ROMs will also provide valuable resource material for teachers and lecturers The CD ROMs are designed for use on a PC running Windows 95 98 ME or 2000

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*An Introduction to Chemical Equilibrium and Kinetics* Louis Meites, 1981

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working chemist's point of view this book aims to prepare chemists to devise experiments to test different hypothesis. A number of examples from research literature have been included. **Chemical News and Journal of Physical Science**, 1892

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