

Worksheet: Writing and Balancing Chemical Reactions

1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.

- a. $\text{Cu}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CuO}_{(s)}$
- b. $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{O}_{2(g)}$
- c. $\text{Fe}_{(s)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{Fe}_3\text{O}_{4(s)}$
- d. $\text{AsCl}_{3(laq)} + \text{H}_2\text{S}_{(aq)} \rightarrow \text{As}_2\text{S}_{3(s)} + \text{HCl}_{(aq)}$
- e. $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}_{(s)} \rightarrow \text{CuSO}_{4(s)} + \text{H}_2\text{O}_{(g)}$
- f. $\text{Fe}_2\text{O}_{3(s)} + \text{H}_2_{(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- g. $\text{CaCO}_{3(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_{2(g)}$
- h. $\text{Fe}_{(s)} + \text{S}_{8(s)} \rightarrow \text{FeS}_{(s)}$
- i. $\text{H}_2\text{S}_{(aq)} + \text{KOH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{K}_2\text{S}_{(aq)}$
- j. $\text{NaCl}_{(s)} \rightarrow \text{Na}_{(s)} + \text{Cl}_{2(g)}$
- k. $\text{Al}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{H}_{2(g)} + \text{Al}_2(\text{SO}_4)_3(aq)$
- l. $\text{H}_3\text{PO}_4(aq) + \text{NH}_4\text{OH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + (\text{NH}_4)_3\text{PO}_4(aq)$
- m. $\text{C}_3\text{H}_8(g) + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- n. $\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_{3(s)}$
- o. $\text{CH}_4(g) + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- p. $\text{K}_2\text{SO}_4(aq) + \text{BaCl}_{2(aq)} \rightarrow \text{KCl}_{(aq)} + \text{BaSO}_{4(s)}$
- q. $\text{C}_5\text{H}_{12(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- r. $\text{Ca}(\text{OH})_2(aq) + \text{NH}_4\text{Cl}_{(aq)} \rightarrow \text{NH}_4\text{OH}_{(aq)} + \text{CaCl}_{2(aq)}$
- s. $\text{V}_2\text{O}_5(s) + \text{Ca}_{(s)} \rightarrow \text{CaO}_{(s)} + \text{V}_{(s)}$
- t. $\text{Na}_{(s)} + \text{ZnI}_{2(aq)} \rightarrow \text{NaI}_{(aq)} + \text{Zn}_{(s)}$
- u. $\text{C}_7\text{H}_6\text{O}_{3(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- v. $\text{Ca}_{(s)} + \text{N}_{2(g)} \rightarrow \text{Ca}_3\text{N}_{2(s)}$
- w. $\text{Fe}_2\text{O}_{3(s)} + \text{H}_2_{(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- x. $\text{C}_{15}\text{H}_{30}(l) + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- y. $\text{BN}_{(s)} + \text{F}_{2(g)} \rightarrow \text{BF}_{3(s)} + \text{N}_{2(g)}$
- z. $\text{C}_{12}\text{H}_{26}(l) + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$

2. Predict the product(s) along with the states, indicate the type of reaction, and balance the following chemical reactions.

- a. A solution of lead (II) nitrate is mixed with a solution of sodium iodide.
- b. Solid zinc sulfide reacts with oxygen in the air.
- c. Liquid butane ($\text{C}_4\text{H}_{10(l)}$) is used as a fuel to ignite a lighter.
- d. Barium hydroxide solution is neutralized by adding hydrochloric acid ($\text{HCl}_{(aq)}$).
- e. Copper metal is placed in a solution of silver nitrate.
- f. Sulfur burns in oxygen to make sulfur dioxide gas.
- g. A solution of aluminum sulfate is mixed with a solution of calcium hydroxide.
- h. Zinc metal is placed in sulfuric acid ($\text{H}_2\text{SO}_{4(aq)}$).
- i. Aluminum powder is placed in a container filled with chlorine gas.
- j. Sucrose undergoes cellular respiration.

Chemical Reactions And Energy Worksheet Answers

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Essential Practices for Managing Chemical Reactivity Hazards Robert W. Johnson, Steven W. Rudy, Stephen D. Unwin, 2010-08-13 In its recent investigation of chemical reactivity accidents the US Chemical Safety Board noted a gap in technical guidance and regulatory coverage This volume closes the gap in technical guidance helping small and large companies alike identify address and manage chemical reactivity hazards It guides the reader through an analysis of the potential for chemical reactivity accidents to help prevent fires explosions toxic chemical releases or chemical spills This volume is applicable to processes at any scale and is particularly useful for chemists safety managers and engineers involved in scale up An enclosed CD ROM provides portable checklists analysis tools and a list of additional references Note CD ROM DVD and other supplementary materials are not included as part of eBook file

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