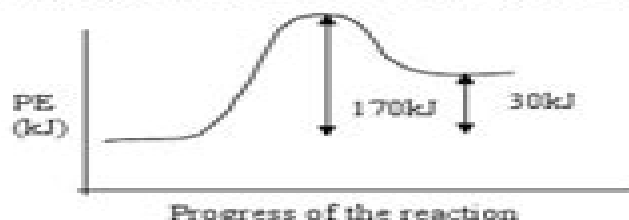


Chapter 1 – Kinetics - Practice Test #1

Name - _____

The answers to this test are posted online

1. Consider the following reaction mechanism:
- $$\begin{array}{lclclcl} \text{step 1:} & \text{M} & + & \text{X} & \longrightarrow & \text{MX} \\ \text{step 2:} & \text{MX} & + & \text{A} & \longrightarrow & \text{D} + \text{X} \end{array}$$
- The chemical species MX is a(n)
- A. catalyst
B. inhibitor
C. final product
D. reaction intermediate
2. Consider the following reaction: $2\text{N}_2\text{O}_{5(g)} \rightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$
At a certain temperature the rate of decomposition of $2\text{N}_2\text{O}_{5(g)}$ is $2.5 \times 10^{-6} \text{ mol/s}$. The rate of formation of NO_2 is
- A. $1.0 \times 10^{-5} \text{ mol/s}$
B. $1.3 \times 10^{-6} \text{ mol/s}$
C. $2.5 \times 10^{-6} \text{ mol/s}$
D. $5.0 \times 10^{-6} \text{ mol/s}$
3. Which of the following factors affect the rates of both homogeneous and heterogeneous reactions:
- I nature of the reactants
II presence of a catalyst
III temperature of system
IV concentration of reactants
- A. I and IV only
B. II and III only
C. II, III, and IV only
D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?
- A. $\text{N}_2\text{O}_{4(g)} + 59 \text{ kJ} \rightarrow 2\text{NO}_{2(g)}$
B. $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)} + 572 \text{ kJ}$
C. $2\text{BrCl}_{(g)} - 29.3 \text{ kJ} \rightarrow \text{Br}_{2(g)} + \text{Cl}_{2(g)}$
D. $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(g)} \quad \Delta H = -572 \text{ kJ}$
5. Consider the potential energy diagram. The activation energy for the reverse reaction is
- A. 30 kJ
B. 140 kJ
C. 170 kJ
D. 200 kJ



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