

# General Chemistry - Unit 1 (Part 1) Review WS - **ANSWERS**

**Directions:** Complete the following questions on a separate piece of paper.

**LT #2: Understand the experiments and the historical theories that lead to the modern atomic theory.**

- Describe Dalton's beliefs about the atom and chemical reactions with atoms.  
*All elements are composed of atoms; all atoms of a given element are identical, that atoms of different elements are different; chemical compounds are formed when atoms combine; atoms are not created or destroyed in chemical reactions (indivisible); solid sphere.*
- Describe the atom according to Thomson's theory.  
*Thomson found that the atom consisted of a positive mass but there were negative pieces scattered throughout the positive area (Plum Pudding Model). He discovered the electron (negatively charged).*
- Describe Rutherford's gold foil experiment. How did he change the model of the atom after the experiment?  
*Rutherford discovered the nucleus contained protons by shooting alpha particles (+) through gold foil. He found that most went through but some bounced back. He found the nucleus, which is very, very small and dense is found in the center of the atom, with the electrons around the outside.*
- What did Chadwick change about the atomic model?  
*Chadwick (along with the Curies) found that the nucleus also contained the neutrons, which have approximately the same mass as the proton, but have no charge.*

**LT #3: Understand how subatomic particles determine the properties of an atom or ion.**

- List the three subatomic particles, their location, their charge and their relative mass.  
*Proton—found in nucleus, + charge, 1 amu (atomic mass unit)  
Neutron—found in nucleus, no charge, 1 amu  
Electron—cloud surrounding the nucleus, - charge, approx. no mass*
- What is a cation and how is it formed?  
*A cation is a positive ion and is formed when an atom loses one or more electrons.*
  - What is an anion and how is it formed?  
*An anion is a negative ion and is formed when an atom gains one or more electrons.*
- Complete the following charts for atoms and ions:

ATOM	ISOTOPE SYMBOL	ATOMIC #	MASS #	#PROTONS	#ELECTRONS	#NEUTRONS
Sodium - 23	<sup>23</sup> <sub>11</sub> Na	11	23	11	11	12
Aluminum - 27	<sup>27</sup> <sub>13</sub> Al	13	27	13	13	14
Chlorine - 35	<sup>35</sup> <sub>17</sub> Cl	17	35	17	17	18
Barium - 136	<sup>136</sup> <sub>56</sub> Ba	56	136	56	56	80

SYMBOL	MASS	#PROTONS	#ELECTRONS	#NEUTRONS	NET CHARGE
K <sup>+1</sup>	39	19	18	20	+1
Br <sup>-1</sup>	80	35	36	45	-1
Fe <sup>+3</sup>	56	26	23	30	+3
Se <sup>-2</sup>	79	34	36	45	-2

- Are all atoms of the same element identical? If not, how can they differ?  
*No; isotopes differ by # of neutrons. (Isotopes have the same # of protons and electrons.)*
- Which of lead's four isotopes is most abundant: lead-204, lead-206, lead-207 or lead-208? Explain how you know which isotope is more abundant. *The average atomic mass for lead is 207.2 amu; therefore, the most abundant isotope is lead-207*

**LT #4: Apply the mole concept (molar mass, mole equalities, conversions/dimensional analysis).**

- Calculate the molar mass of each of the following atoms or compounds:
  - zinc  
*65.4 g/mol*
  - barium phosphate, Ba<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>  
*3(137.3) + 2(31.0) + 8(16.0) = 601.9 g/mol*

# Chemistry Unit 1 Review Answers

**Kevin T. Patton, Frank B. Bell, Terry  
Thompson, Peggie L. Williamson**



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Daily Skill-Builders: Vocabulary 4-5 ,      **Clinical Chemistry** Theresa Dean, Sheryl Whitlock, 1997 This complete procedural manual offers laboratory technicians hands on coverage of all aspects of clinical chemistry Using numerous real world case studies it fully examines test principles and demonstrates correct test techniques and procedures Effective quality control and safety procedures are incorporated throughout 55 illus      **5 Steps to a 5: AP Chemistry 2021** Mary Millhollon, Richard H. Langley, 2020-10-02 MATCHES THE LATEST EXAM In this hybrid year let us supplement your AP classroom experience with this multi platform study guide The immensely popular 5 Steps to a 5 AP Chemistry guide has been updated for the 2020 21 school year and now contains 3 full length practice exams available both in the book and online that reflect the latest exam Up to Date Resources for COVID 19 Exam Disruption Access to a robust online platform Comprehensive overview of the AP Chemistry exam format Hundreds of practice exercises with thorough answer explanations Proven strategies specific to each section of the test A self guided study plan including flashcards games and more online      *Operating Reactors Licensing Actions Summary* , 1983      **Title List of Documents Made Publicly Available** U.S. Nuclear Regulatory Commission, 1982      **South-Western GED Science** Fern Burch, 1995      **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** , 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction      **MCAT Success** , 1999      **Holt Chemistry** Ralph Thomas Myers, 2004      Anatomy & Physiology with Brief Atlas of the Human Body and Quick Guide to the Language of Science and Medicine - E-Book Kevin T. Patton, Frank B. Bell, Terry Thompson, Pegg L. Williamson, 2022-03-21 A P may be complicated but learning it doesn't have to be Anatomy Physiology 11th Edition uses a clear easy to read approach to tell the story of the human body's structure and function Color coded illustrations case studies and Clear View of the Human Body transparencies help you see the Big Picture of A P To jump start learning each unit begins by reviewing what you have already learned and previewing what you are about to learn Short chapters simplify concepts with bite size chunks of information Conversational storytelling writing style breaks down information into brief chapters and chunks of information making it easier to understand concepts 1 400 full color photographs and drawings bring difficult A P concepts to life and illustrate the most current scientific knowledge UNIQUE Clear View of the Human Body transparencies allow you to peel back the layers of the body with a 22 page full color insert showing the male and female human body along several planes The Big Picture and Cycle of Life sections in each chapter help you comprehend the interrelation of body systems and how the structure and function of these change in relation to age and development Interesting sidebars include boxed features

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