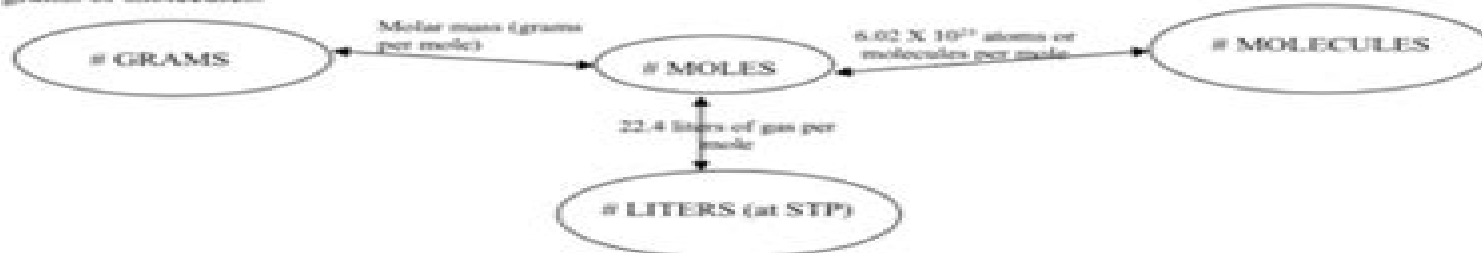


You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP. We can use this information to convert grams to molecules or liters, molecules to grams or liters, or liters to grams or molecules.



EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP?

$$\# \text{ liters Ne} = 40.36 \text{ g Ne} \times \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \times \frac{22.4 \text{ L Ne}}{1 \text{ mole Ne}} = 44.80 \text{ L Ne}$$

EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP?

$$\# \text{ molecules CO}_2 = 56.0 \text{ L CO}_2 \times \frac{1 \text{ mole CO}_2}{22.4 \text{ L CO}_2} \times \frac{6.02 \times 10^{23} \text{ molecules CO}_2}{1 \text{ mole CO}_2} = 1.51 \times 10^{24} \text{ molecules CO}_2$$

SOLVE THE FOLLOWING PROBLEMS ON A SEPARATE SHEET OF PAPER.

- YOU MUST USE COMPLETE AND PROPER SET-UPS.
- SHOW THE MOLAR MASS CALCULATION WHENEVER THE PROBLEM REQUIRES YOU TO DO ONE. INCLUDE UNITS.
- CIRCLE YOUR FINAL ANSWER.

- What would be the volume, in liters, of 85.5 grams of carbon monoxide at STP?
- How many molecules would there be in 0.500 grams of carbon disulfide?
- What would be the mass, in grams, of 45.0 liters of nitrogen at STP?
- How many molecules of hydrogen are in a balloon full of hydrogen with a volume of 5.34 liters at STP?
- Your mommy buys you a helium balloon at the circus. It has a volume of 4.00 liters at STP. What mass of helium, expressed in grams, does this balloon contain?
- How many molecules of ammonia would there be in 40.0 grams of ammonia?
- What would be the mass, in grams, of 3.50×10^{23} molecules of chlorine?
- What volume, expressed in liters, would 50.0 grams of fluorine occupy at STP?
- How many grams of oxygen would there be in 1.00 liter of oxygen at STP?

Extra Challenge:

- How many molecules of water are there in 10 lbs of water?

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