

Nitrogen dioxide, NO₂, can be made by reacting nitrogen monoxide, NO, with oxygen, O₂.



A chemist mixes 0.450 mol NO with 0.450 mol O₂ in a container. This container is then heated and equilibrium is reached.

At equilibrium, the mixture contains 0.400 mol NO and the total pressure is 400 kPa. Calculate K_p and the units

Step 1 - Write out the equation and fill in the initial moles

	2NO	+	O ₂	⇌	2NO ₂
Initial	0.450		0.450		0
Change	-0.05		-0.025		+0.05
Equilibrium	0.400		0.425		0.05

Step 2 - Work out the partial pressures of each gas

$$p(\text{NO}) = 0.400 / 0.875 \times 400 = 182.86 \text{ kPa}$$

$$p(\text{O}_2) = 0.425 / 0.875 \times 400 = 194.29 \text{ kPa}$$

$$p(\text{NO}_2) = 0.05 / 0.875 \times 400 = 22.86 \text{ kPa}$$

Step 3 - Put the values in to the K_p equation and work out the answer

$$K_p = \frac{p(\text{NO}_2)^2}{p(\text{NO})^2 \times p(\text{O}_2)}$$

$$K_p = \frac{(22.86)^2}{(182.86)^2 \times (194.29)}$$

$$K_p = 8.04 \times 10^{-5}$$

Step 4 - Calculate the units

$$K_p = \frac{\cancel{(\text{kPa})^2}}{(\cancel{\text{kPa}})^2 \times (\text{kPa})}$$

$$K_p = \text{kPa}^{-1}$$

Chemistry Chapter 19 Assessment Answers

Richard H. Langley, John Moore



Chemistry Chapter 19 Assessment Answers:

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Table of Contents Chemistry Chapter 19 Assessment Answers

1. Understanding the eBook Chemistry Chapter 19 Assessment Answers
 - The Rise of Digital Reading Chemistry Chapter 19 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 19 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Chapter 19 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 19 Assessment Answers
 - Personalized Recommendations
 - Chemistry Chapter 19 Assessment Answers User Reviews and Ratings

- Chemistry Chapter 19 Assessment Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 19 Assessment Answers Free and Paid eBooks
 - Chemistry Chapter 19 Assessment Answers Public Domain eBooks
 - Chemistry Chapter 19 Assessment Answers eBook Subscription Services
 - Chemistry Chapter 19 Assessment Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 19 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 19 Assessment Answers Compatibility with Devices
 - Chemistry Chapter 19 Assessment Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 19 Assessment Answers
 - Highlighting and Note-Taking Chemistry Chapter 19 Assessment Answers
 - Interactive Elements Chemistry Chapter 19 Assessment Answers
- 8. Staying Engaged with Chemistry Chapter 19 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 19 Assessment Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 19 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 19 Assessment Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 19 Assessment Answers
 - Setting Reading Goals Chemistry Chapter 19 Assessment Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 19 Assessment Answers
 - Fact-Checking eBook Content of Chemistry Chapter 19 Assessment Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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