

# Chemistry

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(Chapter 4) (The d and f block Elements) (Exercise)  
(Class 12)

## Question 4.33:

Compare the chemistry of the actinoids with that of lanthanoids with reference to:

- (i) electronic configuration
- (ii) oxidation states and
- (iii) chemical reactivity.

## Answer 4.33:

- (i) The general electronic configuration for lanthanoids is  $[\text{Xe}]^{54} 4f^{0-14} 5d^{0-1} 6s^2$  and that for actinoids is  $[\text{Rn}]^{86} 5f^{1-14} 6d^{0-1} 7s^2$ . Unlike 4f orbitals, 5f orbitals are not deeply buried and participate in bonding to a greater extent.
- (ii) The principal oxidation state of lanthanoids is (+3). However, sometimes we also encounter oxidation states of +2 and +4. This is because of extra stability of fully-filled and half-filled orbitals. Actinoids exhibit a greater range of oxidation states. This is because the 5f, 6d and 7s levels are of comparable energies. Again, (+3) is the principal oxidation state for actinoids. Actinoids such as lanthanoids have more compounds in +3 state than in +4 state.
- (iii) In the lanthanide series, the earlier members of the series are more reactive. They have reactivity that is comparable to Ca. With an increase in the atomic number, the lanthanides start behaving similar to Al. Actinoids, on the other hand, are highly reactive metals, especially when they are finely divided. When they are added to boiling water, they give a mixture of oxide and hydride. Actinoids combine with most of the non-metals at moderate temperatures. Alkalies have no action on these actinoids. In case of acids, they are slightly affected by nitric acid (because of the formation of a protective oxide layer).

## Question 4.34:

Write the electronic configurations of the elements with the atomic numbers 61, 91, 101, and 109.

## Answer 4.34:

Atomic number	Electronic configuration
61	$[\text{Xe}]^{54} 4f^5 5d^0 6s^2$
91	$[\text{Rn}]^{86} 5f^2 6d^1 7s^2$
101	$[\text{Rn}]^{86} 5f^{12} 5d^0 7s^2$
109	$[\text{Rn}]^{86} 5f^{14} 6d^7 7s^2$

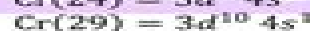
## Question 4.35:

Compare the general characteristics of the first series of the transition metals with those of the second and third series metals in the respective vertical columns. Give special emphasis on the following points:

- (i) electronic configurations, (ii) oxidation states, (iii) ionisation enthalpies, and (iv) atomic sizes.

## Answer 4.35:

- (i) In the 1<sup>st</sup>, 2<sup>nd</sup> and 3<sup>rd</sup> transition series, the 3d, 4d and 5d orbitals are respectively filled. We know that elements in the same vertical column generally have similar electronic configurations. In the first transition series, two elements show unusual electronic configurations:



Similarly, there are exceptions in the second transition series. These are:



# Chemistry Chapter 4 Answers

**Donald J. Wink, Sharon Fetzer-  
Gislason, Sheila McNicholas**



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