

1.2 Mole Conversions Solutions

PRACTICE EXAMPLES- To complete the practice examples (1) write the type of question you have, (2) write the equation with units, (3) show the cancellation of units and (4) calculate the solution.

1. Determine how many atoms are present in 2.5 moles of Silicon? (**15.05×10^{23} atoms**)

MOLES \rightarrow ATOMS
$$\frac{2.5 \text{ _moles}}{1} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ _mole}} = 15.05 \times 10^{23} \text{ atoms}$$

2. How many moles of Silicon is 3.01×10^{23} atoms? (**.5 moles**)

ATOMS \rightarrow MOLES
$$\frac{3.011 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ _mole}}{6.022 \times 10^{23} \text{ atom}} = .5 \text{ moles}$$

3. What is the mass of 4 moles of Iron? (**224 grams**)

MOLES \rightarrow GRAMS
$$\frac{4 \text{ _moles}}{1} \times \frac{55.9 \text{ _grams}}{\text{mole}} = 224 \text{ grams of C}$$

4. How many atoms are in 48.6 grams of Magnesium? (**12.04×10^{23} atoms**)

GRAMS \rightarrow ATOMS
$$\frac{48.6 \text{ _grams}}{1} \times \frac{1 \text{ _mole}}{24.3 \text{ _grams}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ _mole}} \approx 12.04 \times 10^{23} \text{ atoms}$$

5. How many grams are in 18.06×10^{23} atoms of Copper? (**190.5 grams**)

ATOMS \rightarrow GRAMS
$$\frac{18.06 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ _mole}}{6.022 \times 10^{23} \text{ atoms}} \times \frac{63.5 \text{ _grams}}{1 \text{ _mole}} \approx 190.5 \text{ grams}$$

6. How many moles of Nickel is 176 grams? (**3 moles**)

GRAMS \rightarrow MOLES
$$\frac{176 \text{ _grams _of _C}}{1} \times \frac{\text{mole}}{58.7 \text{ _grams}} \approx 3 \text{ moles}$$

Chemistry Conversion Practice Problems With Answers

Morris Hein



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. Chemistry Workbook For Dummies with Online Practice Chris Hren, Peter J.

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technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry An integrated approach is employed in which related general chemistry organic chemistry and biochemistry topics are presented in adjacent chapters This approach helps students see the strong connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change

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Programmed Problem-solving for First-year Chemistry Arnold B. Loebel, 1983

A Learning Program for Chemistry Peter Demmin, 1993-11

Merrill Chemistry

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Prentice Hall Chemistry, 2000

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Solving Problems in Chemistry Robert C. Smoot, Gary K.

Himes, 1983

Understanding Chemistry Fred M. Dewey, 1994

General Chemistry, Reactions First Kevin

Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease

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Chemistry and Chemical Reactivity Mary L. Kotz, John C.

Kotz, 1991

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Student Study

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Chemistry John W. Moore, Conrad

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