

Using the Combined Gas Law

We believe in physics, so we think that the atmosphere is a gas. When the balloon floats up, it is because the air pressure is lower than the pressure inside the balloon. As the balloon rises, it expands. The air inside the balloon expands, and it cools down. By measuring the pressure and the volume of the air inside the balloon, we can determine the temperature of the air inside the balloon. This is the combined gas law.

The Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Where:

P_1 = initial pressure

V_1 = initial volume

T_1 = initial temperature (K)

P_2 = final pressure

V_2 = final volume

T_2 = final temperature (K)

This equation can be used to find the temperature of a gas if the pressure and volume are known. It can also be used to find the pressure of a gas if the volume and temperature are known.

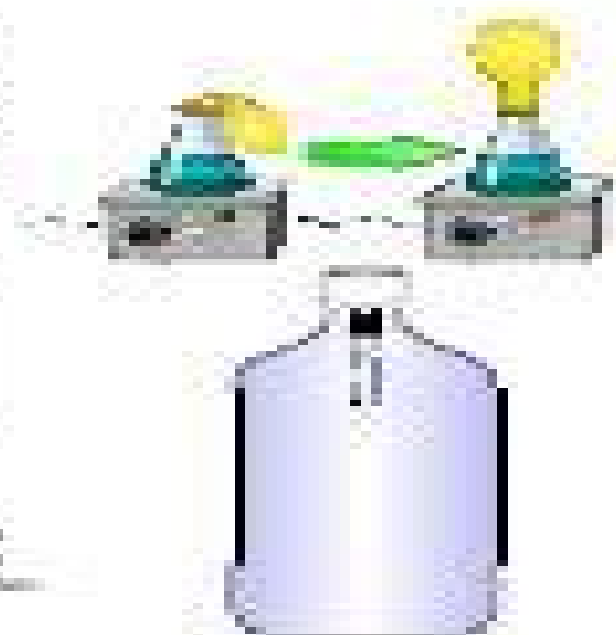
Example Problem

A gas with a volume of 100 mL at 100 kPa and 300 K is heated to 400 K and the pressure is increased to 150 kPa. What is the new volume?

$$P_1 = 100 \text{ kPa}, T_1 = 300 \text{ K}, P_2 = 150 \text{ kPa}, T_2 = 400 \text{ K}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$V_2 = \frac{P_1 V_1 T_2}{P_2 T_1} = \frac{(100 \text{ kPa})(100 \text{ mL})(400 \text{ K})}{(150 \text{ kPa})(300 \text{ K})} = 88.9 \text{ mL}$$



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