

CHEMISTRY PRACTICAL VIVA QUESTIONS

1. What is the principle of volumetric analysis ?

Ans. In volumetric analysis, the concentration of a solution is determined by allowing a known volume of this to react quantitatively with another solution of known concentration.

2. What is titration ?

Ans. The process of adding one solution from the burette to another in the titration flask in order to complete the chemical reaction involved, is known as titration.

3. What is indicator ?

Ans. Indicator is a chemical substance which changes colour at the end point.

4. What is end point ?

Ans. The stage during titration at which the reaction is just complete is known as the end point of titration.

5. Why a titration flask should not be rinsed ?

Ans. This is because during rinsing some liquid will remain sticking to the titration flask therefore the volume taken in the titration flask will increase.

6. What are primary and secondary standard substances ?

Ans. A substance is known as primary standard if it is- available in high degree of purity, if it is stable and unaffected by air, if it does not gain or lose moisture in air, if it is readily soluble and its solution in water remains as such for long time. On the other hand, a substance which does not possess the above characteristics is called a secondary standard substance. Primary standards are crystalline oxalic acid, anhydrous Na_2CO_3 , Mohr's salt, etc.

7. Burettes must be rinsed with the solution with which they are filled, why ?

Ans. The burettes are rinsed with the solution with which they are filled in order to remove any substance sticking to their sides, which otherwise would decrease the volume of the liquids to be taken in them.

8. It is customary to read lower meniscus in case of colourless and transparent solutions and upper meniscus in case of highly coloured solutions, why ?

Ans. Because it is easy to read the lower meniscus in case of colourless solutions, while the upper meniscus in case of coloured solutions.

9. Which indicator is used in the titration of sodium carbonate against hydrochloric acid and what is the colour change at the end point ?

Ans. Methyl orange. The colour change is yellow to pinkish red.

10. What indicator is used in the titration of oxalic acid with sodium hydroxide ? What is the end point ?

Ans. Phenolphthalein. Appearance of pink colour is the end point.

11. What is a standard solution?

Ans. A solution of known concentration is called a standard solution.

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Christian Drosten



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