Period

Atomic Structure

		For	mulas and Con		
$c = \lambda \upsilon$	E = ho	$E = \frac{hc}{\lambda}$	$E_n = -\frac{Rhc}{n^2}$	$\lambda = \frac{h}{mv}$	$\frac{1}{\lambda} = R \left(\frac{1}{2^2} - \frac{1}{n^2} \right)$
c = 3.00 x	108 m/s h	$n = 6.626 \times 1$	0'34 J·s Rhc =	2.18 x 10 ⁻¹⁸ J	$R = 1.0974 \times 10^7 \text{ m}^{-1}$

 Sketch the electron energy levels (n=1 through n=5) for the hydrogen atom.



2. Calculate the energy of an electron in the n=2 energy level of hydrogen. Calculate the energy of an electron in the n=3 energy level. What is the difference in energy of these two levels? If a photon of light had this energy, what would its wavelength be? E1 = - 2.18 410 15 = -2.42 410 15

FL= -2.1841015 = -5.45410 193 Use the Rydberg equation above to calculate the wavelength of a photon when n=3. How does this compare with your answer in question 2?

1 = 1.0974410° m (= - 1) \ \lambda = (6.56410 - m)

 An electron moves from the n=5 to the n=1 quantum level and emits a photon with an energy of 2.093 x 10-18 J. How much energy must the atom absorb to move an electron from n=1 to n=5? What is y= F= (0850400-1427) (5.668 4108 W/2) the wavelength of this energy? The stime decement 2043410-103

= 9.5 ×10-8 m = 19.491 ×10-8 m

 An electron moves with a velocity of 2.5 x 10⁸ cm/s. What is its wavelength? (The mass of an electron is 9.109 x 10⁻²⁸ g.)

20 h = 6.626405 34 3.1 = 2.9 ×10 0 m)

 Calculate the wavelength (in nanometers) associated with a 1.0 x 10²-g golf ball moving at 30. m/s (about 67 mph). How fast must the ball travel to have a wavelength of 5.6 x 10⁻³ nm?

7- h = 6.626 + 10-34 5.5 = [2.2 ×10-34 m] = [2.2 ×10-34 m] = [2.2 ×10-34 m]

(56410-12m) 109m 5.6 ×10 3 = 6626 ×10 34 5-3

Chapter 21 Nuclear Chemistry Mixed Review Answers

Thomas Griffiths

Chapter 21 Nuclear Chemistry Mixed Review Answers

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