

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ <b>lattice energy</b> _____	<b>a.</b> ionic bond
_____ <b>bond energy</b> _____	<b>b.</b> covalent bond
_____ <b>enthalpy of vaporization</b> _____	<b>c.</b> metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ <b>ionic</b> _____	<b>a.</b> H—F	_____ <b>nonpolar covalent</b> _____	<b>d.</b> H—H
_____ <b>ionic</b> _____	<b>b.</b> Na—Cl	_____ <b>polar covalent</b> _____	<b>e.</b> H—C
_____ <b>polar covalent</b> _____	<b>c.</b> H—O	_____ <b>polar covalent</b> _____	<b>f.</b> H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# Chapter 8 Covalent Bonding Answers Key

**Arshad Iqbal**



## Chapter 8 Covalent Bonding Answers Key:

*Chemistry for the Utterly Confused* John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject From formulas and lab techniques to the periodic table *Chemistry for the Utterly Confused* focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts This invaluable guide also teaches problem solving skills you need to master this imposing subject Whether you re in high school in college or simply brushing up on chemistry knowledge this fun easily accessible book will make understanding chemistry a breeze **A Level Chemistry MCQ (Multiple Choice Questions)** Arshad Iqbal, 2019-06-18 The A Level Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF A Level Chemistry MCQ PDF Download Quiz Questions Chapter 1 28 Practice Tests with Answer Key IGCSE GCE Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs A Level Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests A Level Chemistry MCQ PDF book helps to practice test questions from exam prep notes The A Level Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs A Level Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters Alcohols and esters atomic structure and theory benzene chemical compound carbonyl compounds carboxylic acids acyl compounds chemical bonding chemistry of life electrode potential electrons in atoms enthalpy change equilibrium group IV groups II and VII halogenoalkanes hydrocarbons introduction to organic chemistry ionic equilibria lattice energy moles and equations nitrogen and sulfur organic and nitrogen compounds periodicity polymerization rates of reaction reaction kinetics redox reactions and electrolysis states of matter transition elements tests for college and university revision guide A Level Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book IGCSE GCE Chemistry MCQs Chapter 1 28 PDF includes high school question papers to review practice tests for exams A Level Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for IGCSE NEET MCAT GRE GMAT SAT ACT competitive exam A Level Chemistry Mock Tests Chapter 1 28 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Alcohols and Esters MCQ Chapter 2 Atomic Structure and Theory MCQ Chapter 3 Benzene Chemical Compound MCQ Chapter 4 Carbonyl Compounds MCQ Chapter 5 Carboxylic Acids and Acyl Compounds MCQ Chapter 6 Chemical Bonding MCQ Chapter 7 Chemistry of Life MCQ Chapter 8 Electrode Potential MCQ Chapter 9 Electrons in Atoms MCQ Chapter 10 Enthalpy Change MCQ Chapter 11 Equilibrium MCQ Chapter 12 Group IV MCQ Chapter 13 Groups II and VII MCQ Chapter 14 Halogenoalkanes MCQ Chapter 15 Hydrocarbons MCQ Chapter 16 Introduction to Organic Chemistry MCQ Chapter 17 Ionic Equilibria MCQ Chapter 18 Lattice Energy MCQ Chapter 19 Moles and Equations MCQ Chapter 20 Nitrogen and Sulfur MCQ Chapter 21 Organic and Nitrogen Compounds MCQ Chapter 22

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## Chapter 8 Covalent Bonding Answers Key Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the ability of words has be more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Chapter 8 Covalent Bonding Answers Key**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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