

IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



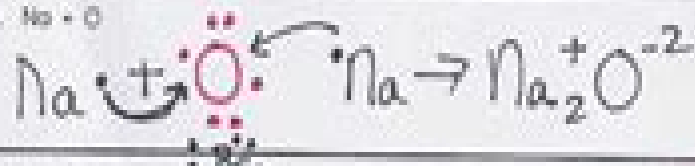
2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:



Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



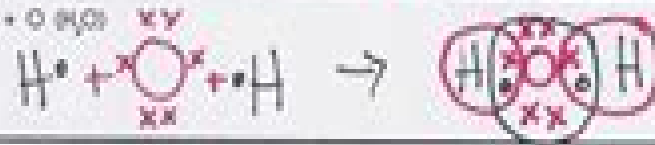
4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO})$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers

Joseph A. Mascetta



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Table of Contents Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers

1. Understanding the eBook Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Personalized Recommendations
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers User Reviews and Ratings
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers and Bestseller Lists

5. Accessing Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers Free and Paid eBooks
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers Public Domain eBooks
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers eBook Subscription Services
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers Budget-Friendly Options
6. Navigating Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers Compatibility with Devices
 - Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Interactive Elements Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
8. Staying Engaged with Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Setting Reading Goals Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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