



COVALENT BONDING

Practice Problems

Do your best to do better than your previous problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2SO_4 b. CH_3NH_2 c. HCl
 d. H_2O e. NaOH

2. Write the formula for the following molecular compounds.

- a. carbon tetrachloride b. phosphorus pentachloride
 c. sulfur dioxide d. dinitrogen pentoxide

3. Write the Lewis structure for each of the following molecular compounds.
 4. Write the Lewis structure for each of the following molecular compounds.
 5. Write the Lewis structure for each of the following molecular compounds.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .

4. Draw the Lewis structure for acetylene, C_2H_2 .



5. Draw the Lewis structure for the following compound.



SECTION 8.3 BONDING: THE ORBITAL

1. Write the Lewis structure for the following compound.
 2. Write the Lewis structure for the following compound.

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