

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- ☒ c. transition metal
- d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

Chapter 9 Chemical Names And Formulas Section Review Answers

Ralph Thomas Myers



Chapter 9 Chemical Names And Formulas Section Review Answers:

Understanding Chemistry Fred M. Dewey, 1994 *Essential Concepts of Chemistry* Sharon Sherman, Alan Sherman, 1999
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Addison-Wesley Chemistry, 1990 **Essentials of Chemistry** Dennis D. Staley, Antony C. Wilbraham, 1984 *Prentice Hall Chemistry*, 2000 Introductory Chemistry Victor S. Krimsley, Darold E. Skerritt, Beverly B. Harrison, 1986 This pedagogically rich text has all the necessary features to hook introductory students and keep them interested and successful in preparatory chemistry This book carefully guides beginning students through the fundamental principals and calculations required for their subsequent success in either the general chemistry course or the short organic and biochemistry course Krimsley while dedicated to conceptual understanding and skill building presents a solid book that provides students with complete explanations on every point and helps them work through the material methodically with many examples and hints His approach is cohesive and coherent always patient and interesting Krimsley introduces all topics through an example students are already familiar with He continually reminds them of objectives and provides many opportunities to practice and then check their answers The text begins with a study of atomic and molecular structure prior to treating the various categories of chemical reactions The organization is designed to get students into chemistry quickly yet methodically The classification of matter begins in Chapter 2 before the chapter on measurement and the coverage of bonding appears in Chapter 8 The elementary concepts of chemistry are presented with an emphasis on mathematical calculations and problem solving strategies **Holt Chemistry** Ralph Thomas Myers, 2004 **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 *Foundations of College Chemistry* 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity *Fundamentals of Chemistry* Ralph A. Burns, 2003 For one semester preparatory chemistry

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