

Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
 - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
 - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
 - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{1 \text{ mol CaO}} \cdot \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$. Remember the ratios can be reduced as well!!!
 - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}}$
- $\frac{165.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{165.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \cdot \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} = \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \cdot \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{108.99 \text{ g CaCO}_3}$
- Molar Mass
 - Definition: the mass of one mole of a substance
 - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen $\rightarrow$ Water
 - 5.0 mol H₂
 - 2.5 mol O₂
- Ethane combustion reaction
 - 15.8 mol O₂
 - 9.00 mol CO₂ & 13.5 mol H₂O
- 575 g Na & 886 g Cl₂
- Iron production reaction
 - 75.1 mol CO
 - 50.1 mol Fe & 75.1 mol CO₂
- 87,390 g CO or $8.739 \times 10^5 \text{ g CO}$ & 12,610 g H₂ or $1.261 \times 10^4 \text{ g H}_2$
- Lightning reaction of N₂ & O₂
 - 1,100 g NO₂ or $1.10 \times 10^3 \text{ g NO}_2$
 720. g NO or $7.20 \times 10^2 \text{ g NO}$
- CO₂ removal in spacecraft
 - 42.04 mol NaOH
 - 21.02 mol Na₂CO₃ & 21.02 mol H₂O
- Double replacement AgNO₃ & NaBr
 - 463 g NaBr
 - 845 g AgBr
- Soda-acid fire extinguisher
 - 3.058 mol NaHCO₃
 - 3.058 mol CO₂, 1.529 mol Na₂SO₄, 3.058 mol H₂O
- Sulfuric acid-sodium hydroxide neutralization reaction
 - H₂SO₄ + 2 NaOH $\rightarrow$ Na₂SO₄ + 2 H₂O
 - 37 g H₂SO₄
 - 53 g Na₂SO₄ & 14 g H₂O
- Copper & silver nitrate reaction
 - 0.0104 mol Cu(NO₃)₂
 - 0.0104 mol Cu & 0.0209 mol AgNO₃
- Aspirin production
 - 13.5 kg C₆H₅O₄
 - 7.66 kg C₆H₅O₃
 - 4.29 L HC₂H₃O₂
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

Chapter 9 Stoichiometry Answers

Yi-Tong Ma



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Table of Contents Chapter 9 Stoichiometry Answers

1. Understanding the eBook Chapter 9 Stoichiometry Answers
 - The Rise of Digital Reading Chapter 9 Stoichiometry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Stoichiometry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Stoichiometry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Stoichiometry Answers
 - Personalized Recommendations
 - Chapter 9 Stoichiometry Answers User Reviews and Ratings
 - Chapter 9 Stoichiometry Answers and Bestseller Lists
5. Accessing Chapter 9 Stoichiometry Answers Free and Paid eBooks
 - Chapter 9 Stoichiometry Answers Public Domain eBooks
 - Chapter 9 Stoichiometry Answers eBook Subscription Services
 - Chapter 9 Stoichiometry Answers Budget-Friendly Options

6. Navigating Chapter 9 Stoichiometry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Stoichiometry Answers Compatibility with Devices
 - Chapter 9 Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Stoichiometry Answers
 - Highlighting and Note-Taking Chapter 9 Stoichiometry Answers
 - Interactive Elements Chapter 9 Stoichiometry Answers
8. Staying Engaged with Chapter 9 Stoichiometry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Stoichiometry Answers
9. Balancing eBooks and Physical Books Chapter 9 Stoichiometry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Stoichiometry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Stoichiometry Answers
 - Setting Reading Goals Chapter 9 Stoichiometry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Stoichiometry Answers
 - Fact-Checking eBook Content of Chapter 9 Stoichiometry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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