

1. A 33.7-g sample of HF is dissolved in water to give 2.0×10^2 mL of solution. The concentration of the solution is:

A) 1.68 M
B) 0.34 M
C) 0.17 M
D) 8.4 M
E) 16.9 M

$$[HF] = \frac{33.7g HF}{2.0 \times 10^2 mL} \times \frac{1000 mL}{1 L} \times \frac{1 mol HF}{20.01g HF} = 8.4 M$$

2. What mass of calcium chloride, $CaCl_2$, is needed to prepare 3.419 L of a 1.56 M solution?

A) 243 g
B) 5.33 g
C) 50.6 g
D) 111 g
E) 592 g

$$\frac{3.419 L}{1} \times \frac{1.56 mol CaCl_2}{L} \times \frac{110.98g CaCl_2}{1 mol CaCl_2} = 592g CaCl_2$$

3. A 51.24-g sample of $Ba(OH)_2$ is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar $Ba(OH)_2$?

A) 403 mL
B) 248 mL
C) 24.8 mL
D) 83.0 mL
E) 331 mL

$$[Ba(OH)_2] = \frac{51.24g Ba(OH)_2}{1.205 L} \times \frac{1 mol Ba(OH)_2}{171.36g Ba(OH)_2} = 0.2481 M$$

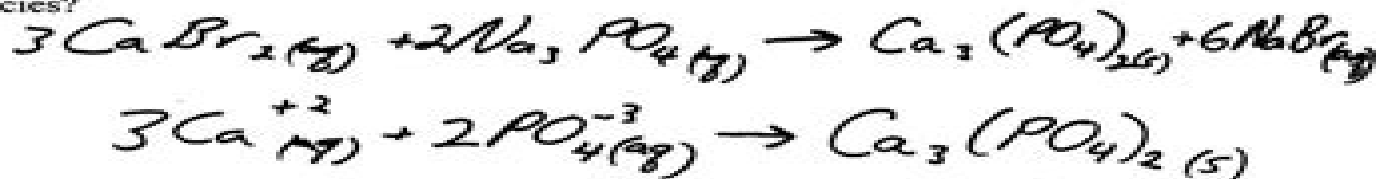
$$M_1 = 0.1000 M \quad M_2 = 0.2481 M$$

$$V_1 = 1.000 L \quad V_2 = ?$$

$$V_2 = \frac{0.1000 M \times 1.000 L}{0.2481 M} = 0.4031 L = 403 mL$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

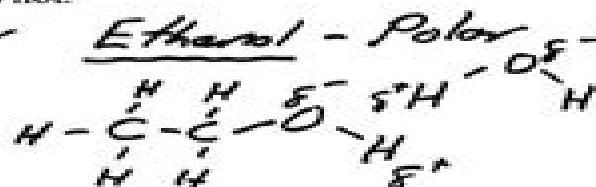
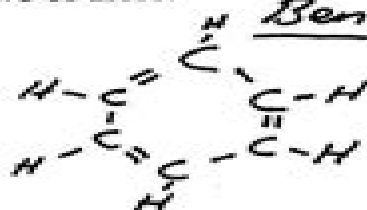
A) $Ca^{2+}(aq)$
B) $PO_4^{3-}(aq)$
C) $2Ca_3(PO_4)_2(s)$
D) $6NaBr(aq)$
E) $3Ca^{2+}(aq)$



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

A) They have different molar masses.
B) One is ionic, the other is not.
C) One is an electrolyte, the other is not.
D) Ethanol contains a polar O—H bond, and benzene does not.
E) Two of these.

* Know the general solubility rule of "Likes Dissolve Likes"



Chapter 4 Chemistry Test

Carolyn A. Sink, Nicole M. Weinstein



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Table of Contents Chapter 4 Chemistry Test

1. Understanding the eBook Chapter 4 Chemistry Test
 - The Rise of Digital Reading Chapter 4 Chemistry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Chemistry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Chemistry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Chemistry Test
 - Personalized Recommendations
 - Chapter 4 Chemistry Test User Reviews and Ratings
 - Chapter 4 Chemistry Test and Bestseller Lists

5. Accessing Chapter 4 Chemistry Test Free and Paid eBooks
 - Chapter 4 Chemistry Test Public Domain eBooks
 - Chapter 4 Chemistry Test eBook Subscription Services
 - Chapter 4 Chemistry Test Budget-Friendly Options
6. Navigating Chapter 4 Chemistry Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Chemistry Test Compatibility with Devices
 - Chapter 4 Chemistry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Chemistry Test
 - Highlighting and Note-Taking Chapter 4 Chemistry Test
 - Interactive Elements Chapter 4 Chemistry Test
8. Staying Engaged with Chapter 4 Chemistry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Chemistry Test
9. Balancing eBooks and Physical Books Chapter 4 Chemistry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Chemistry Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Chemistry Test
 - Setting Reading Goals Chapter 4 Chemistry Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Chemistry Test
 - Fact-Checking eBook Content of Chapter 4 Chemistry Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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