

CHAPTER 5 REVIEW*The Periodic Law***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. c In the modern periodic table, elements are ordered
 - (a) according to decreasing atomic mass.
 - (b) according to Mendeleev's original design.
 - (c) according to increasing atomic number.
 - (d) based on when they were discovered.
2. d Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing
 - (a) density.
 - (b) reactivity.
 - (c) atomic number.
 - (d) atomic mass.
3. b The modern periodic law states that
 - (a) no two electrons with the same spin can be found in the same place in an atom.
 - (b) the physical and chemical properties of an element are functions of its atomic number.
 - (c) electrons exhibit properties of both particles and waves.
 - (d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
4. c The discovery of the noble gases changed Mendeleev's periodic table by adding a new
 - (a) period.
 - (b) series.
 - (c) group.
 - (d) level.
5. d The most distinctive property of the noble gases is that they are
 - (a) metallic.
 - (b) radioactive.
 - (c) metalloid.
 - (d) largely unreactive.
6. c Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4.
 - (b) 10.
 - (c) 11.
 - (d) 18.
7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.
 - 9 a. How many protons are in this atom?
 - 10 b. How many neutrons are in this atom?
 - ${}^{19}_{9}\text{F}$ c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

Chapter 5 Review The Periodic Law Answers Section 3

Rachel S Tattersall



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