

## Chemical Bonding

### SECTION 3

**SHORT ANSWER** Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one  
 (a) formula unit. (c) crystal.  
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of  
 (a) molecules. (c) dipoles.  
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is  
 (a) higher in potential energy. (c) equal in potential energy.  
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?  
 (a) Compound A is not an ionic compound.  
 (b) It will be more difficult to break the bonds in compound A than those in compound B.  
 (c) Compound B has larger crystals than compound A.  
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally  
 (a) stronger than the attractive forces among formula units in ionic bonding.  
 (b) weaker than the attractive forces among formula units in ionic bonding.  
 (c) approximately equal to the attractive forces among formula units in ionic bonding.  
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

- What type of energy best represents the strength of an ionic bond?

**Lattice energy**

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Chapter 6 Chemistry Chemical Bonding

**Wismer**



## Chapter 6 Chemistry Chemical Bonding:

A-level Chemistry E. N. Ramsden, 2000 Each topic is treated from the beginning without assuming prior knowledge Each chapter starts with an opening section covering an application These help students to understand the relevance of the topic they are motivational and they make the text more accessible to the majority of students Concept Maps have been added which together with Summaries throughout aid understanding of main ideas and connections between topics Margin points highlight key points making the text more accessible for learning and revision Checkpoints in each chapter test students understanding and support their private study Topological Approaches to the Chemical Bond Ángel Martín Pendás, Julia Contreras-García, 2023-02-22 This graduate textbook provides comprehensive information on topological analysis in real space of the electronic structure Application of the topological tools is becoming routine for understanding the outcome of quantum chemical calculations This title thoroughly reviews a selection of currently available topological tools their use and spectrum of applications and provides graduate students and researchers with information not easily obtained from the available textbooks The book is accompanied by worked examples exercises and solutions and is a great tool for any quantum chemistry or computational chemistry course at the graduate and advanced undergraduate levels **The Chemical Bond** Ahmed Zewail, 1992-05-14 This inspired book by some of the most influential scientists of our time including six Nobel laureates chronicles our emerging understanding of the chemical bond through the last nine decades and into the future From Pauling's early structural work using x ray and electron diffraction to Zewail's femtosecond lasers that probe molecular dynamics in real time from Crick's molecular biology to Rich's molecular recognition this book explores a rich tradition of scientific heritage and accomplishment The perspectives given by Pauling Perutz Rich Crick Porter Polanyi Herschbach Zewail and Bernstein celebrate major scientific achievements in chemistry and biology with the chemical bond playing a fundamental role In a unique presentation that also provides some lively insights into the very nature of scientific thought and discovery The Chemical Bond Structure and Dynamics will be of general interest to scientists science historians and the scientifically inclined populous **The Chemical Bond** Tadamas Shida, 2013-03-09 Providing the quantum mechanical foundations of chemical bonding this unique textbook emphasizes key concepts such as superposition degeneracy of states and the role of the electron spin These quantum mechanical notions are usually oversimplified or meticulously circumvented in other books to the frustration of serious readers who want to understand for example why two protons can be stably bound with only one electron to make the simplest molecule  $H_2$  An initial concise and compact presentation of the rudiments of quantum mechanics enables readers to progress through the book with a firm grounding Experimental examples are included to illustrate how the abstract concepts are manifest in real systems Cchemistry Textbook for College and University USA Ibrahim Sikder, 2023-06-04 Cchemistry Textbook USA Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps

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