

### Types of Chemical Reactions Worksheet #1

#### H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
- 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
- 3)  $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$  **Double Replacement**
- 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
- \*5)  $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
- 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
- 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
- 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
- 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**
- \*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**

\*Remember: "H<sub>2</sub>O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition:  $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement:  $3\text{Ca(OH)}_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe(OH)}_3 + 3\text{CaCl}_2$
- 16) Single replacement:  $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$  [Assume Fe<sup>2+</sup>]
- 17) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$  [Copper (II) is used here]

# Chapter 7 Chemical Reactions Answer Key

**Robert R. Redfield**



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## **Table of Contents Chapter 7 Chemical Reactions Answer Key**

1. Understanding the eBook Chapter 7 Chemical Reactions Answer Key
  - The Rise of Digital Reading Chapter 7 Chemical Reactions Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Chemical Reactions Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Chemical Reactions Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Chemical Reactions Answer Key
  - Personalized Recommendations
  - Chapter 7 Chemical Reactions Answer Key User Reviews and Ratings
  - Chapter 7 Chemical Reactions Answer Key and Bestseller Lists

5. Accessing Chapter 7 Chemical Reactions Answer Key Free and Paid eBooks
  - Chapter 7 Chemical Reactions Answer Key Public Domain eBooks
  - Chapter 7 Chemical Reactions Answer Key eBook Subscription Services
  - Chapter 7 Chemical Reactions Answer Key Budget-Friendly Options
6. Navigating Chapter 7 Chemical Reactions Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 7 Chemical Reactions Answer Key Compatibility with Devices
  - Chapter 7 Chemical Reactions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 7 Chemical Reactions Answer Key
  - Highlighting and Note-Taking Chapter 7 Chemical Reactions Answer Key
  - Interactive Elements Chapter 7 Chemical Reactions Answer Key
8. Staying Engaged with Chapter 7 Chemical Reactions Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 7 Chemical Reactions Answer Key
9. Balancing eBooks and Physical Books Chapter 7 Chemical Reactions Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 7 Chemical Reactions Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 7 Chemical Reactions Answer Key
  - Setting Reading Goals Chapter 7 Chemical Reactions Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 7 Chemical Reactions Answer Key
  - Fact-Checking eBook Content of Chapter 7 Chemical Reactions Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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