

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfate, the roman numeral tells us (2)
 - How many atoms of the anion are in the formula unit.
 - How many sulfate ions can be attached to the iron atom.
 - The charge on each Fe ion.
 - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is to (4)
 - Change the number of moles represented in the formula.
 - Change the charges on the other ions in the compound.
 - Change the formula so that it no longer represents the compound.
 - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{HC}_2\text{H}_3\text{O}_2$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II chloride	PbCl_2
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Hydrochloric acid	HCl
Dinitrogen pentoxide	N_2O_5
Copper II carbonate	CuCO_3
Sodium sulfate	Na_2SO_4
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Barium sulfate	BaSO_4
Nitrous acid	HNO_2
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Sulfur trioxide	SO_3
Sulfur dioxide	SO_2
Barium perchlorate	$\text{Ba}(\text{ClO}_4)_2$
Tetraphosphorus trisulfide	P_4S_3
Phosphoric acid	H_3PO_4

- Name the following compound and identify its molecular formula (acid)
 - $\text{Ca}(\text{HCO}_3)_2$ calcium bicarbonate - II
 - $\text{Fe}(\text{NO}_3)_3$ iron III nitrate - II
 - $\text{Fe}(\text{ClO}_4)_3$ iron III perchlorate - II
 - N_2O_5 dinitrogen pentoxide - III

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K Morrison



Chapter 9 Chemical Names Formulas Answers Page 221:

High School Chemistry Michael A. Buben, 2018-05-18 The Teacher's Edition is designed to add direction to High School Chemistry which is an outline of notes. The books go hand in hand. Modified Second Edition as a student's workbook and guidance from the Teacher's Edition. Basic Chemical Principles Donada Peters, Edward I. Peters, 1988. General, Organic, and Biological Chemistry John R. Amend, Bradford P. Mundy, Melvin T. Arnold, 1990. Very Good. No Highlights or Markup. All pages are intact. **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11. Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day. Relevant content, lively explorations, and a wealth of hands-on activities take students' understanding of science beyond the page and into the world around them. Now includes even more technology tools and activities to support differentiated instruction. **Journal of Chemical Education**, 1929. **Chemistry: The Study of Matter** Henry Dorin, 1982. **The Wiswesser Line-formula Chemical Notation (WLN)** Elbert G. Smith, Peter A. Baker, 1975. **American Professional Pharmacist**, 1947. **Chemistry: Core Concepts, 3rd Edition** Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09. The third edition of Chemistry Core Concepts Blackman et al. has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry. Available as a full colour printed textbook with an interactive eBook code, this title enables every student to master concepts and succeed in assessment. Lecturers are supported with an extensive and easy to use teaching and learning package. **Addison-Wesley Chemistry**, 1990. The Chemical News and Journal of Physical Science, 1863. **Chemical News and Journal of Physical Science**, 1863. *Select Thermodynamic Models for Process Simulation* Jean-Charles de Hemptinne, Jean-Marie Ledanois, 2012. The selection of the most adequate thermodynamic model in a process simulation is an issue that most process engineers have to face sooner or later. This book, conceived as a practical guide, aims at providing adequate answers by analysing the questions to be looked at. The analysis in the first chapter yields three keys that are further discussed in three different chapters: 1. A good understanding of the properties required in the process and their method of calculation is the first key. The second chapter provides to that end, in a synthetic manner, the most important equations that are derived from the fundamental principles of thermodynamics. 2. An adequate description of the mixture, which is a combination of models and parameters, is the second key. The third chapter makes the link between components and models, both from a numerical parameterisation and physical molecular interactions point of view. Finally, 3. a correct view of the phase behaviour and trends in regard of the process conditions is the third key. The fourth chapter illustrates the phase behaviour and makes model recommendations for the most significant industrial systems. A decision tree is provided at the end of this chapter. In the last chapter, the key questions are reviewed for a number of typical processes. This book is intended for process engineers who

are not specialists of thermodynamics but are confronted with this kind of problems and need a reference book as well as process engineering students who will find an original approach to thermodynamics complementary of traditional lectures

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