

Acids, Bases & Salts

Important Terms

Acids- Substances that dissociate in water to give H^+ or H_3O^+ ions are called acids [according to Arrhenius concept].

Bases- Substances that dissociate in water to give OH^- (hydroxyl) ions are called bases [according to Arrhenius concept].

Neutralization reaction:

Acids + Bases

→ Salt + Water



water-soluble
bases = alkalis

pH represents the strength of an acid/base in terms of Hydrogen ion concentration

Strong acid + Strong Base = Neutral Salt + Water

Strong Acid + Weak Base = Acidic Salt + Water

Weak Acid + Strong Base = Basic Salt + Water

Conjugate base
acid - ion/radical
left after either
1) H^+ ion (proton) is
released after acid
dissociates
2) ion/radical
left after base
accepts proton (H^+
ion)

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Adams



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The Enigmatic Realm of **Chem 12 Notes On Acids Bases Sss Chemistry**: Unleashing the Language is Inner Magic

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