

9/10

Name _____

Date _____

11-3-22

Calorimetry and Hess's Law

Prelab

1. Zinc metal reacts with hydrochloric acid according to the balanced equation:



When 0.103 g of Zn (s) is combined with enough HCl to make ^{60.0 g} 50.0 mL of solution in a coffee-cup calorimeter, all of the zinc reacts, raising the temperature of the solution from 22.5°C to 23.7°C. Find ΔH_{rxn} for this reaction as written. Use 1.0 g/mL for the density of the solution and 4.18 J/(g·°C) as the specific heat capacity. Ignore the heat absorbed by the calorimeter.

$$\Delta T = 23.7^\circ\text{C} - 22.5^\circ\text{C} = 1.2^\circ\text{C}$$

$$q_{\text{rxn}} = - (M_w)(c_w)(\Delta T_w)$$

$$q_{\text{rxn}} = - (50.0 \text{ g})(4.18 \text{ J/g}^\circ\text{C})(1.2^\circ\text{C})$$

$$q_{\text{rxn}} = -250.8 \text{ J} = -1000 \text{ J}$$

$$0.103 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.38 \text{ g/mol Zn}} = 0.0015754 \text{ mol Zn}$$

$$\Delta H_{\text{rxn}} = \frac{-250.8 \text{ J}}{0.0015754 \text{ mol Zn}}$$

$$\Delta H_{\text{rxn}} = -159.2 \text{ kJ/mol}$$

2. Calculate ΔH_{rxn} for the reaction:



Use the following reactions and given ΔH 's.



$$\Delta H_{\text{rxn}} = +74.6 + -95.7 + -369.3$$

$$\Delta H_{\text{rxn}} = -280.4 \text{ kJ}$$

Chem Hess Law Lab Answer

Patrick Vollmar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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