

Types of Chemical Reactions Worksheet #1

11. Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
 - 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
 - 3) $\text{HClOH} + \text{AlCl}_3 \rightarrow \text{AlOHCl}_2 + 2\text{HCl}$ **Double Replacement**
 - 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
 - *5) $\text{CaOHCl}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
 - 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
 - 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
 - 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
 - 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
 - *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**
- *Remember: " H_2O " can also be thought of as " HOH "

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 1.1) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 1.2) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 1.3) Decomposition: $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 1.4) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 1.5) Double replacement: $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 1.6) Single replacement: $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$ [Assume Fe^{2+}]
- 1.7) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 1.8) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 1.9) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 2.0) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$ [Copper (II) is used here]

Chemfax Lab Answers Chemical Reactions

Daniel R. Faust



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Fizz, Gurgle, Pop! Chemical Reactions in the Lab Daniel R. Faust, 2020 Whether they are testing a new drug creating new materials or checking the safety of our food and water all scientists follow the scientific method One of the major steps in the scientific method is testing your hypothesis That s when scientists hit the lab This volume stresses the important of safety in the laboratory whether it s a school science lab or the most high tech research facility It also examines the proper rules and procedures everyone must follow when working with chemicals in the lab Fact boxes and full color photographs support the clear age appropriate text *Chemical Reactions and Their Equations* Ingo Waldemar Dagobert Hackh, 1928

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Practice Makes Perfect Chemistry Review and Workbook, Second Edition Marian DeWane, Heather Hattori, 2018-12-28

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Chemical Reactions Anne O'Daly, 2026 This book looks at chemical reactions the processes in which the atoms of substances rearrange to make different substances **Encyclopedia of Chemical Reactions** C. A. Jacobson, 1946

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