

# M O C K T E S T P A P E R

# 1

## Section-II Chemistry (Domain)

Maximum Marks: 200

Time : 60 Minutes

### General Instructions:

- This paper consists of 50 MCQs, attempt any 40 out of 50.
- Correct answer or the most appropriate answer: Five marks (+5)
- Any incorrect option marked will be given minus one mark (-1).
- Unanswered/Marked for Review will be given no mark (0).
- If more than one option is found to be correct then Five marks (+5) will be awarded to only those who have marked any of the correct options.
- If all options are found to be correct then Five marks (+5) will be awarded to all those who have attempted the question.
- If none of the options is found correct, or a question is found to be wrong, or a question is dropped, then all candidates who have attempted the dropped question will be given five marks (+5) marks.

1. 2-Bromopentane, 2-Bromo-2-methylbutane, 1-Bromopentane

Write the compound which is most reactive towards  $\beta$ -elimination reaction.

- 2-Bromopentane
  - 1-Bromopentane
  - 2-Bromo-2-methylbutane
  - None of the above
2. In calcium fluoride, having the fluorite structure, the coordination number for calcium ion ( $\text{Ca}^{2+}$ ) and fluoride ion ( $\text{F}^-$ ) are:
- 4 and 2
  - 6 and 6
  - 8 and 4
  - 4 and 8
3. Which of the following reactions is an example of auto-reduction?
- $\text{Fe}_2\text{O}_3 + 4\text{CO} \rightarrow 3\text{Fe} + 4\text{CO}_2$
  - $\text{Cu}_2\text{O} + \text{C} \rightarrow 2\text{Cu} + \text{CO}$
  - $\text{Cu}^{2+}(\text{aq.}) + \text{Fe}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Fe}^{2+}(\text{aq.})$
  - $\text{Cu}_2\text{O} + \frac{1}{2}\text{Cu}_2\text{S} \rightarrow 3\text{Cu} + \frac{1}{2}\text{SO}_2$
4. In the following sequence of reactions
- $$\text{CH}_3 - \text{Br} \xrightarrow{\text{KCN}} \text{A} \xrightarrow{\text{H}_3\text{O}^+} \text{B} \xrightarrow[\text{ether}]{\text{LiAlH}_4} \text{C}$$
- the end product (C) is
- acetone
  - methane
  - acetaldehyde
  - ethyl alcohol
5. A brown ring is formed in the ring test for  $\text{NO}_3^-$  ion. It is due to the formation of:

- $[\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{2+}$
- $\text{FeSO}_4 \cdot \text{NO}_2$
- $[\text{Fe}(\text{H}_2\text{O})_4(\text{NO})_2]^{2+}$
- $\text{FeSO}_4 \cdot \text{HNO}_3$

6. The  $d$ -electron configurations of  $\text{Cr}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$  and  $\text{Co}^{2+}$  are  $d^4$ ,  $d^5$ ,  $d^6$  and  $d^7$  respectively. Which one of the following will exhibit minimum paramagnetic behaviour?

(At. no. Cr = 24, Mn = 25, Fe = 26, Co = 27)

- $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$
- $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$
- $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$
- $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$

7. The reaction in which the aqueous solution of sodium salt of carboxylic acids on electrolysis give alkanes:

- Soda lime decarboxylation
- Kolbe's electrolysis decarboxylation
- Dry distillation of calcium formate
- Reduction of carboxylic acid.

8. Structure of a mixed oxide is cubic close packed (ccp). The cubic unit cell of mixed oxide is composed of oxide ions. One fourth of the tetrahedral voids are occupied by divalent metal A and the octahedral voids are occupied by a monovalent metal B. The formula of the oxide is :

- $\text{ABO}_2$
- $\text{A}_2\text{BO}_2$
- $\text{A}_2\text{B}_3\text{O}_4$
- $\text{AB}_2\text{O}_2$

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**Chao Zhang**



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