

### C. Properties of solution, colloids, and suspensions

<u>Solutions</u>	<u>Colloids</u>	<u>Suspensions</u>
<u>Homogeneous</u>	Heterogeneous	Heterogeneous
Particle size: 0.01-1 nm; can be atoms, ions, molecules	Particle size: 1-1000 nm, dispersed; can be aggregates or large molecules	Particle size: over 1000 nm, suspended; can be large particles or aggregates
<u>Do not</u> separate on standing	<u>Do not</u> separate on standing	Particles settle out
Cannot be separated by filtration	Cannot be separated by filtration	<u>Can</u> be separated by filtration
<u>Do not</u> scatter light	Scatter light (Tyndall effect)	<u>May</u> scatter light, but are not transparent

### D. Determining if a mixture is a true solution, a colloid, or a suspension:



- If particles settle or can be filtered out = suspension
- If particles *DO NOT* settle or filter out shine a beam of light (Tyndall effect) through the mixture
  - If the Tyndall effect is observed = colloid
  - If the Tyndall effect is *NOT* observed = solution

# Chemistry 21a Experiment 6 Solutions And Colloids

## Introduction

**VICTOR. POSCHL**

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**Applied Colloid and Surface Chemistry** Richard M. Pashley, Marilyn E. Karaman, 2004-12-10. Applied Colloid and Surface Chemistry is a broad introduction to this interdisciplinary field. Taking a genuinely applied approach with applications drawn from a wide range of industries, this book will meet the demands of the student and professional currently working in the field. The text includes keynote sections written by practicing industrial research scientists bringing to the reader a wealth of real industrial examples. These examples range from water treatment through to soil management as well as examples taken from the coatings and photographic industries. To aid accessibility, some of the more demanding mathematical derivations are separated from the main text, enabling them to be avoided as required. With carefully structured chapters starting with learning objectives and containing tutorial questions with answers and explanatory notes, this text is invaluable for undergraduates taking a first course on colloid and surface chemistry. This book will also be suitable to postgraduates and professionals who need an up to date account of the subject.      **Introduction to Colloid Chemistry**

Rarol J. Mysels, 1967      *Introduction to Colloid Science* W. J. Papiel, 1978-01-01      *Colloid Chemistry* Jerome Alexander, 1926      *Applied Colloid Chemistry* Wilder Dwight Bancroft, 1932. Adsorption of gas or vapor by solid. Chemical reactions. Adsorption of vapor by liquid and of liquid and solid by solid. Adsorption from solution. Surface tension. Brownian movements. Coalescence. Preparation of colloidal solutions. Electrical properties of colloidal solutions. Stability of colloidal solutions. Gelatinous precipitates and jellies. Emulsions and foams. Non aqueous colloidal solutions. Fog. Smoke. Gases and solids in solids. Thickness of surface films.      **Introduction to Colloid and Surface Chemistry** Duncan James Shaw, 2000

**Introduction to Colloid and Surface Chemistry** Duncan J. Shaw, 1966      **Introduction to Colloid Chemistry** Karol Joseph Mysels, 1950      Laboratory Manual of Colloid Chemistry Harry Nicholls Holmes, 2023-07-18. This laboratory manual provides a detailed look at the field of colloid chemistry. It includes experiments and procedures for studying the properties of colloidal solutions as well as an overview of the theory behind these solutions. Holmes's writing is clear and concise, and the manual is an invaluable resource for anyone studying chemistry or materials science. This work has been selected by scholars

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