

## Information: Orbitals

So far we have learned that inside energy levels there are different sublevels. Now we will look at orbitals. **Orbitals** are located inside sublevels just like streets are located inside subdivisions. Different sublevels have different numbers of orbitals.

TABLE 2

Sublevel	# of Orbitals Possible
s	1
p	3
d	5
f	7

Here's an important fact: only two electrons can fit in each **orbital**. So, in an s orbital you can have a maximum of 2 electrons; in a d orbital you can have a maximum of 2 electrons; in any **orbital** there can only be two electrons.

Since a d **sublevel** has 5 orbitals (and each orbital can contain up to two electrons) then a d **sublevel** can contain 10 electrons ( $= 5 \times 2$ ). Pay attention to the difference between "sublevel" and "orbital".

## Critical Thinking Questions

- How many orbitals are there in a p sublevel? 3
- How many orbitals are there in a d sublevel? 5
- How many total **sublevels** would be found in the entire 2<sup>nd</sup> energy level? Two, since the 2<sup>nd</sup> energy level has s and p sublevels.
  - How many **orbitals** would be found in the entire 2<sup>nd</sup> energy level? 4 orbitals; the 2s sublevel has 1 orbital and the 2p sublevel has 3 orbitals for a total of 4 orbitals.
- How many electrons can fit in an f **sublevel**?  
(2 electrons per orbital)  $\times$  (7 orbitals in an f sublevel) = 14 electrons.
  - How many electrons can fit in an f **orbital**?  
Any **orbital** can only hold 2 electrons.
- How many electrons can fit in a d orbital? in a p orbital? in any kind of orbital?  
Each **orbital** can hold a maximum of 2 electrons.
- In your own words, what is the difference between a sublevel and an orbital?  
Orbitals are contained within sublevels. A sublevel is a grouping of orbitals.
- How many electrons can fit in each of the following **energy levels**:  
 1<sup>st</sup> energy level = 2 because only 2 can fit in an s sublevel and the first energy level only has an s sublevel.  
 2<sup>nd</sup> energy level = 8; 2 in the 2s sublevel and 6 in the 2p sublevel giving a total of 8.  
 3<sup>rd</sup> energy level = 18; 2 in the 3s sublevel, 6 in the 3p sublevel, and 10 in the 3d sublevel giving a total of 18.  
 4<sup>th</sup> energy level = 32; 2 in the 4s, 6 in the 4p, 10 in the 4d, and 14 in the 4f.

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