

Lab: Baking Soda Stoichiometry

Introduction: When baking soda (NaHCO_3) is an ingredient in your recipe, its purpose is to make the batter rise and produce a product with a light and fluffy texture. The decomposition reaction is shown below:



Our goal will be to predict how much sodium carbonate (Na_2CO_3) is produced when baking soda decomposes.

Materials: [Lab Video \(use to fill in data table\)](#)

Procedure:

1. Use the lab video to complete the data table below.
2. Calculate the mass of the NaHCO_3 used in the reaction and the experimental mass of Na_2CO_3 that was produced. **You need to show work in the "calculation" table for how you got your answer. Typing it out is okay for the calculation table section.**
3. Answer the 2 conclusion questions at the end.

Data table

Mass of crucible (g)	12.74 g
Mass of crucible and NaHCO_3 (g)	14.33 g
Mass of crucible and Na_2CO_3 (g)	13.75 g

Calculations Table (show work in each box)

Mass of NaHCO_3 (g)	1.59 g
Experimental (actual) mass of Na_2CO_3 (g)	1.01 g

Conclusion Questions: Show all work, handwritten, using the correct steps/methods shown during class. You will need to upload an image of your work.

1. Use your starting mass of NaHCO_3 and the balanced chemical equation above to calculate the mass of Na_2CO_3 that **should** be produced by the decomposition reaction. This is the theoretical yield.

Baking Soda Stoichiometry Lab Answers

Jennifer Louise Pakkala



Baking Soda Stoichiometry Lab Answers:

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Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007

The Science Teacher , 2005 SCC Library has 1964 cur

Molecular Level Drawings as a Road Map for Conclusion Writing in a High School Classroom Patrick Mackey Bohman, 2008

Holt Chemistry Ralph Thomas Myers, 2004

Subject Guide to Books in Print , 1997

A Stoichiometry Unit David Callaghan, 2004

Stoichiometry Unit Project Luann Marie Decker, 1998

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