

### Basic Stoichiometry Post-Lab Homework Exercises

1. Load the "Reactants, Products, and Leftovers" simulation and work through each of the levels of the Game! At <http://phet.colorado.edu/> or googling "phet." You may have to download or update Java on your computer.

**Show your work for all calculations. Remember: Labels, units, sig figs and cancellation!**

2. For the reaction  $P_4 + 6Cl_2 \rightarrow 4PCl_3$ , determine how many moles of chlorine  $Cl_2$  would be needed to react with 4.0 moles of phosphorus  $P_4$  to entirely use up all the phosphorus. 4)\_\_\_\_\_

3. If 3.00 moles of  $P_4$  reacted with 22.0 moles  $Cl_2$  according to the above reaction, determine:

- a) How many moles  $PCl_3$  are produced a)\_\_\_\_\_
- b) How many moles of  $P_4$  are left in excess after the reaction (if any) b)\_\_\_\_\_
- c) How many moles of  $Cl_2$  are left in excess after the reaction (if any) c)\_\_\_\_\_

In reality, reactants don't have to react in perfect whole-numbers of moles. In a two-reactant synthesis reaction, usually **one reactant gets entirely used up (and determines how much product is made)**, even if that means using fractions of a mole of reactant. For instance, when solid, metallic aluminum Al and red, liquid bromine  $Br_2$  are brought together, they make a white solid according to the reaction  $2Al + 3Br_2 \rightarrow 2AlBr_3$ . If 5.0 moles of aluminum Al was reacted with 10 moles bromine  $Br_2$ , all five moles of aluminum would react, with only 7.5 moles bromine. (2:3 mole ratio) This would produce only 5.0 moles of  $AlBr_3$ , leaving 2.5 moles of excess  $Br_2$  behind.

5. Now assume 3.0 moles Al and 4.0 moles  $Br_2$  react.

- a) How much (in moles)  $AlBr_3$  gets produced? a)\_\_\_\_\_
- b) Which chemical must be the *excess reactant*? b)\_\_\_\_\_
- c) Which chemical is the *limiting reactant*? c)\_\_\_\_\_
- d) If all the limiting reactant gets used up, how much of the excess reactant is left? d)\_\_\_\_\_

# Basic Stoichiometry Post Lab Homework Exercises Answers

**RS Peters**



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