

THERMOCHEMISTRY

THERMOCHEMICAL EQUATIONS

Chemical reactions such as explosions and the burning of fuels are exothermic which means energy is released. The amount of energy released depends on the amount of explosive or fuel used in the reaction that is, energy released is directly proportional to the amount of reactant or reactants used.

When water evaporates, energy is absorbed from the surroundings (endothermic reaction) as the state of water changes from liquid to a gas. Once again, the amount of energy absorbed depends on the amount of water that evaporates.

Chemical reactions involving the release or the absorption of energy, are represented by thermochemical equations which have the following features:

1. Chemical formulae of reactants and products.
2. Coefficients of amounts in moles for reactants and products.
3. States of reactants and products.
4. Amount of energy either released or absorbed when a specified amount of reactants is used.

EXAMPLE 1

Combustion of methane gas:



The above equation is telling us that when 1 mole of methane gas reacts with 2 mole of oxygen, 890 kJ of energy are released. If 2 mole of methane were to react, then 4 mole of oxygen would be required and the amount of energy released would be 1780 kJ.

EXAMPLE 2

Photosynthesis:



This means that when 6 mole of carbon dioxide react with 6 mole of water, 1 mole of glucose is formed and 2803 kJ of energy are absorbed. If 2 mole of glucose were to be produced, 12 mole of carbon dioxide would react with 12 mole of water and 5606 kJ of energy would be absorbed.

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Guillaume Favre



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