The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks) pH = 4.65 so [H30+] = antilog (-4.65) = 2.239x10 M

11	0.25	0	0
rt.	7 339/165	+2.239×10-5	+2.239×10"
, 1	2 +37/10 5	2.239×10-5	2.239×10-7

Ka=	[H30+][BrO]
	CHB-0]
-	(2.239×10-5) = 2.0×10-9
T <sub>a</sub>	nssser Ka = 2.0 × 10-9

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) pH = 2.355 [Hat]: antilog (-2.355) = 0.004416 M

Co	0	0
-0.004416	+0.004416	+0.004416
Co-0.004416	0.004416	0.004416

= 0.30439 M Find the value of Kb for the oxalate ion (C<sub>2</sub>O<sub>4</sub><sup>2</sup>). (1 mark)

$$K_{b}(c_{2}c_{4}^{2-}) = \frac{Kw}{Ka(Hc_{2}c_{4}^{-})} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

Answer Kb = 1.6×10-10

 The value of Kb for the weak base methylamine (CH<sub>1</sub>NH<sub>2</sub>) is 4.4 x 10<sup>-4</sup>. Calculate the value of Ka for the acid CH3NH3\*, (1 mark)

$$\int Ka_{(CH_3NH_3)} = \frac{Kw}{K_b(CH_3NH_2)}.$$

$$Ka_{1}(CH_{3}NH_{3}^{+}) = \frac{Kw}{k_{b}(CH_{3}NH_{z})} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

Answer  $Ka = 2.3 \times 10^{-11}$ 

# **Basic Stoichiometry Phet Lab Answers**

**JR Anderson** 

Basic Stoichiometry Phet Lab Answers:

The book delves into Basic Stoichiometry Phet Lab Answers. Basic Stoichiometry Phet Lab Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Basic Stoichiometry Phet Lab Answers, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Basic Stoichiometry Phet Lab Answers
  - Chapter 2: Essential Elements of Basic Stoichiometry Phet Lab Answers
  - Chapter 3: Basic Stoichiometry Phet Lab Answers in Everyday Life
  - $\circ\,$  Chapter 4: Basic Stoichiometry Phet Lab Answers in Specific Contexts
  - ∘ Chapter 5: Conclusion
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