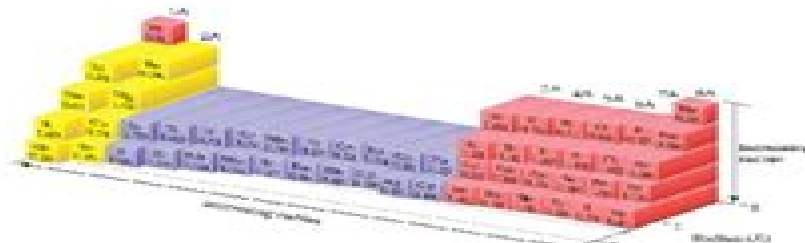


Chapter 7 Assignment



1- Which part of the periodic table (top/bottom, left/right) has the elements with the largest atoms?

- a- Bottom and left
- b. Bottom and right
- c. Top and left
- d. Top and right

2- Why is it easier to remove a $2p$ electron from an oxygen atom than from a nitrogen atom?



- a. The $2p$ electron removed from an oxygen atom is in a fully occupied orbital with decreased electron–electron repulsions compared to the one from a nitrogen atom.
- b. The $2p$ electron removed from an oxygen atom is in a more than half-filled subshell, which has extra stability compared to the one from a nitrogen atom.
- c. The $2p$ electron removed from an oxygen atom is in a fully occupied orbital with increased electron–electron repulsions compared to the one from a nitrogen atom.

3- How do the periodic trends in metallic character compare to those for ionization energy?

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