

Topic 1: Quantitative Chemistry

1.1 Introduction to the particulate nature of matter and chemical change

- 1.1.1 Atoms of different elements combine in fixed ratios to form compounds, which have different properties from their component elements
- 1.1.2 Mixtures contain more than one element and/or compound that are not chemically bonded together and so retain their individual properties
- 1.1.3 Mixtures are either homogeneous or heterogeneous
- 1.1.4 Deduction of chemical equations when reactants and products are specified
- 1.1.5 Application of the state symbols (s), (l), (g), and (aq) in equations
- 1.1.6 Explanation of observable changes in physical properties and temperatures during changes of state

Particle Nature of Matter

- Matter is anything that takes up space
- Matter can either refer to the particles (pure substances) or combination of a substances (mixtures):

Pure Substances

- A pure substance has definite and constant composition
- For a pure substance, from a particle perspective all particles will look and remain the same

Element – Atoms all having the same number of protons

Molecule – Two or more elements chemically join together

Compounds – Two or more different elements chemically joined together in a fixed ratio

- From their definitions: All compounds are molecules, but not all molecules are compounds.
- When the elements are joined, the atoms lose their individual properties and have different properties from the elements they are composed of

Mixtures

- Mixture: A combination of pure substances
- Mixtures contain more than one element and/or compound that are not chemically bonded together, so retain their individual properties
- Mixtures are either homogeneous or heterogeneous:
- Homogeneous mixtures are the same mixture throughout
 - They will have a uniform composition
- Heterogeneous mixtures have a different mixture throughout
 - They will have visibly different substances or phases throughout, a non-uniform composition



Chemical Equation

- Chemical Equation: Describes what happens during a chemical reaction
- A chemical reaction will always have reactants and products as well as some special reaction conditions if required
- Reactants are always on the left, and products are always on the right:

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