# B. Titration of the Acid of unknown concentration with the Standardized NaOH

- Record your unknown ID.
- 2. Obtain less than 10mL of an unknown acid in a clean, dry 50-mL beaker.
- Rinse the 10.00-mL pipet with three 2-mL portions of your unknown solution.
   Obtain ~ 40mL of unknown acid and pipet 10.00mL of it into the rinsed 250-mL flask.
- Repeat the titration procedure above using the acid of unknown concentration in place of the standard acid.

Set up a data table to organize your data for the H2SO4/NaOH titration.

	Trial 1	Trial 2	Trial 3
Unknown ID			
V <sub>H2SO4</sub> (mL)	10.00	10.00	10.00
Avg Mouson (M), from Part A		10.00	10.00
Viritial NaOH (mL)			
Virus, NaCH (mL)			
Δ V <sub>NaOH</sub> (mL)			
MH2SO4 (M)			

#### Calculations

In these titrations, one mole of HCl reacts with one mole of NaOH, and one mole of  $H_2SO_4$  reacts with two moles of NaOH. The molarity, M, of the acid or the base is expressed in moles per liter.

In part A the molarity of the HCl solution is known. The molarity of NaOH may be calculated from the results of part A through the following stoichiometry reflecting the overall equation:

- M<sub>NaOH</sub>, the molarity of the NaOH calculated from the data. Show one calculation
  of the NaOH molarity in detail
- 2. Average value of the three determinations of NaOH molarity.
- Percent range of the three determinations:

In part B the molarity of the NaOH solution determined from part A is used to compute the molar concentration of the H<sub>2</sub>SO<sub>4</sub> in the solution of unknown concentration:

$$H_2SO_4(aq)$$
 + 2 NaOH(aq)  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub>(aq) + 2 H<sub>2</sub>O(I)

- MH2SO4, the molarity of the H2SO4 calculated from the data. Show one calculation of the H2SO4 molarity in detail
- Average value of the three determinations of H₂SO₄ molarity.
- Percent range of the three determinations:

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**ML Yell** 

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