

Student Name: _____

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CHEM 2218 Inorganic Chemistry I (Final Exam sample paper)

Part I: Multiple Choices (2 points each, 50 points in total)

- For a cube, the number of C_6 , C_4 , and C_2 rotational axes is
(a) 1, 3, and 4, respectively (b) 0, 2, and 4, respectively (c) 0, 3, and 4, respectively
- The CO diatomic molecule contains infinite number of
(a) σ_v (b) i (c) σ_h (d) C_2
- Identify the wrong label(s) in the following list used for representations of point groups.
 C_{1000} , C_{100} , D_{1000} , D_{2v} , C_4 , I , O_h , T_d
- The $Fe_2(CO)_9$ molecule shown does not possess what element of symmetry?
(a) i (b) σ (c) S_6 (d) C_2
- A molecule having a point group of D_{6h} contains all the symmetry elements of a molecule having a point group of
(a) D_{3d} (b) D_{2d} (c) D_{3h} (d) none of the above
- How many mirror planes of symmetry are there in an icosahedron?
(a) 6 (b) 9 (c) 12 (d) 15 (e) 18
- Give the symmetry label for the molecular orbital associated with a trigonal planar structure (*you are given a character table for D_{3h}*).
(a) A_1^+ (b) A_1^-
(c) A_2^+ (d) A_2^-
- Identify the stronger acid?
 H_2Se vs. H_2S $[Al(H_2O)_6]^{3+}$ vs. $[Fe(H_2O)_6]^{3+}$
- Which one in the following is the least stable?
(a) $[Al(H_2O)_6]^{3+}$ (b) $[AlF_6]^{3-}$ (c) $Mo(H_2O)_6$ (d) $Mo(CO)_6$
- Predict which way the following reactions will go (left or right)
(a) $[AlF_6]^{3-} + Ga(CN)_6]^{3-} \rightarrow [GaF_6]^{3-} + Al(CN)_6]^{3-}$
(a) $W(CO)_9(PMe_3) + NMe_3 \rightarrow W(CO)_9(NMe_3) + PMe_3$
- Select the strongest base toward proton.
(a) NH_3 (b) $NHMe_2$ (c) PH_3 (d) PMe_3
- Which of the following reactions belong to the acid-base reactions?
(a) $H + OH \rightarrow H_2O$.
(b) $PF_3 + F_2 \rightarrow PF_5$.
(c) $SiF_4 + 2NaF \rightarrow Na_2[SiF_6]$.
(d) $Mo + 6CO \rightarrow Mo(CO)_6$.



Acs Inorganic Chemistry Exam

Terry C. Jones



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